



Quantitation of 11 alkylamines in atmospheric samples: separating structural isomers by ion chromatography

Bryan K. Place^{1,2}, Aleya T. Quilty², Robert A. Di Lorenzo¹, Susan E. Ziegler², and Trevor C. VandenBoer²

¹Department of Chemistry, Memorial University, St. John's, NL, Canada

²Department of Earth Sciences, Memorial University, St. John's, NL, Canada

Correspondence to: Trevor C. VandenBoer (tvandenboer@mun.ca)

Received: 11 October 2016 – Discussion started: 15 November 2016

Revised: 23 February 2017 – Accepted: 24 February 2017 – Published: 16 March 2017

Abstract. Amines are important drivers in particle formation and growth, which have implications for Earth's climate. In this work, we developed an ion chromatographic (IC) method using sample cation-exchange preconcentration for separating and quantifying the nine most abundant atmospheric alkylamines (monomethylamine (MMAH⁺), dimethylamine (DMAH⁺), trimethylamine (TMAH⁺), monoethylamine (MEA⁺), diethylamine (DEAH⁺), triethylamine (TEAH⁺), monopropylamine (MPAH⁺), isomonopropylamine (iMPAH⁺), and monobutylamine (MBAH⁺)) and two alkyl diamines (1, 4-diaminobutane (DABH⁺) and 1, 5-diaminopentane (DAPH⁺)). Further, the developed method separates the suite of amines from five common atmospheric inorganic cations (Na⁺, NH₄⁺, K⁺, Mg²⁺, Ca²⁺). All 16 cations are greater than 95 % baseline resolved and elute in a runtime of 35 min. This paper describes the first successful separation of DEAH⁺ and TMAH⁺ by IC and achieves separation between three sets of structural isomers, providing specificity not possible by mass spectrometry. The method detection limits for the alkylamines are in the picogram per injection range and the method precision ($\pm 1\sigma$) analyzed over 3 months was within 16 % for all the cations. The performance of the IC method for atmospheric application was tested with biomass-burning (BB) particle extracts collected from two forest fire plumes in Canada. In extracts of a size-resolved BB sample from an aged plume, we detected and quantified MMAH⁺, DMAH⁺, TMAH⁺, MEAH⁺, DEAH⁺, and TEAH⁺ in the presence of Na⁺, NH₄⁺, and K⁺ at molar ratios of amine to inorganic cation ranging from 1 : 2 to 1 : 1000. Quantities of DEAH⁺ and DMAH⁺ of 0.2–200 and 3–1200 ng m⁻³, respectively, were present in the extracts and an unprecedented amine-to-

ammonium molar ratio greater than 1 was observed in particles with diameters spanning 56–180 nm. Extracts of respirable fine-mode particles (PM_{2.5}) from a summer forest fire in British Columbia in 2015 were found to contain iMPAH⁺, TMAH⁺, DEAH⁺ and TEAH⁺ at molar ratios of 1 : 300 with the dominant cations. The amine-to-ammonium ratio in a time series of samples never exceeded 0.15 during the sampling of the plume. These results and an amine standard addition demonstrate the robustness and sensitivity of the developed method when applied to the complex matrix of BB particle samples. The detection of multiple alkylamines in the analyzed BB samples indicates that this speciation and quantitation approach can be used to constrain BB emission estimates and the biogeochemical cycling of these reduced nitrogen species.

1 Introduction

Particles in the atmosphere can modulate climate through their direct and indirect effect on the radiative balance of Earth's atmosphere (Boucher et al., 2013; Lohmann and Feichter, 2005). This potential warming or cooling effect of particles represents the greatest uncertainty in Earth's radiative forcing (Myhre et al., 2013). Additionally, particles with a diameter of 2.5 μm or less (PM_{2.5}) have been classified as carcinogens (IARC, 2016) and are estimated to be responsible for 3 million deaths annually worldwide (Stephen et al., 2012). Thus, understanding the quantities and the chemical and physical nature of the species involved in the formation and growth of new particles is of paramount importance.

Recent work has shown that organic compounds may contribute considerably to particle nucleation (Ehn et al., 2014; Ortega et al., 2016; Tröstl et al., 2016; Willis et al., 2016). In particular, the need to measure and quantify gaseous atmospheric alkylamines has gained interest because of their exceptional ability to partake in atmospheric particle formation. Multiple laboratory investigations have shown the nucleation potential of methyl- and ethyl-substituted amines through gaseous acid–base chemistry reactions (Almeida et al., 2013; Angelino et al., 2001; Berndt et al., 2010, 2014; Bzdek et al., 2010, 2011; Erupe et al., 2011; Jen et al., 2016a, b; Lloyd et al., 2009; Murphy et al., 2007; Qiu et al., 2011; Silva et al., 2008; Smith et al., 2010; Wang et al., 2010a, b; Yu et al., 2012; Zhao et al., 2011; Zollner et al., 2012). Theoretical calculations and studies have also found that amines have a high disposition to form atmospheric nanoparticles (Barsanti et al., 2009; Kurtén et al., 2008; Loukonen et al., 2014, 2010; Nadykto et al., 2015; Ortega et al., 2012). From these works, alkylamines have been shown to form clusters via neutralization reactions at rates up to 3 orders of magnitude greater than ammonia (Almeida et al., 2013; Berndt et al., 2010; Bzdek et al., 2011; Kurtén et al., 2008; Loukonen et al., 2010; Nadykto et al., 2015) and readily exchange with ammonium in ammonium–bisulfate molecular clusters (Bzdek et al., 2010; Lloyd et al., 2009; Qiu et al., 2011). These studies suggest that alkylamines can compete with ammonia to form particles even though they have been quantified at mixing ratios that are 3 or more orders of magnitude lower in the atmosphere (Chang et al., 2003; Ge et al., 2011; Schade and Crutzen, 1995). Atmospheric measurements made during new particle formation events have further confirmed that alkylamines participate in particle formation at ambient concentrations and that these species may be present in most atmospheric particles (Creamean et al., 2011; Dall’Osto et al., 2012; Hodshire et al., 2016; Kulmala et al., 2013; Kürten et al., 2016; Ruiz-Jimenez et al., 2012; Smith et al., 2010; Tao et al., 2015).

Alkylamine emissions to the atmosphere arise from both natural and anthropogenic sources (Ge et al., 2011). Short-chain alkylamines such as the methylated and ethylated amines are predominantly reported in emission estimates. Measurements show that atmospheric alkylamines are prevalent in ambient air across the globe, especially in the particle phase (Ge et al., 2011). For example, methyl- and ethylamines were measured by an aerosol time-of-flight mass spectrometer at both rural and urban sites all across Europe (Healy et al., 2015). In particular, these amines have been measured in substantial quantities near animal husbandry operations (Kuhn et al., 2011; Lunn and Van de Vyver, 1977; Rabaud et al., 2003; Schade and Crutzen, 1995; Sorooshian et al., 2008), fisheries (Seo et al., 2011), and sewage-waste treatment facilities (Leach et al., 1999). Other anthropogenic sources include tobacco smoke (Schmeltz and Hoffmann, 1977), automobiles (Cadle and Mulawa, 1980), and cooking (Rogge et al., 1991; Schauer et al., 1999). The ocean is esti-

mated to be the largest natural source of alkylamines, where they are released as volatile degradation products (Ge et al., 2011; Gibb et al., 1999a, b). Aliphatic amines have also been detected in smoldering stage biomass-burning (BB) plumes. These have been estimated to represent a quarter of global methylated amine emissions (Lobert et al., 1990; Schade and Crutzen, 1995).

Real-time in situ speciation and quantitation of atmospheric amines in the particle and gas phase can be difficult because alkylamines are commonly found at or below parts per trillion by volume (pptv) mixing ratios in the atmosphere (Ge et al., 2011). Furthermore, the atmospheric matrix can be complex and ubiquitous atmospheric species can cause matrix effects for various analytical methods targeting these reduced nitrogen species. Being able to chromatographically resolve alkylamines from the dominant base, ammonium, represents a major challenge when sampling the gas phase (Chang et al., 2003; Ge et al., 2011; Schade and Crutzen, 1995). Quantifying amines in particle samples, for example by ion chromatography (IC), presents a greater challenge due to possible interferences from sodium, potassium, ammonium, magnesium, and calcium whose concentrations are dependent on the particle source characteristics and the measurement location (Ault et al., 2013; Kovac et al., 2013; Sobanska et al., 2012; Sun et al., 2006). Particles frequently contain complex organic mixtures, such as high molecular weight organic compounds, which can cause further matrix effects during separation or direct analysis (Di Lorenzo and Young, 2016; Saleh et al., 2014).

Achieving full speciation of alkylamines is important because the nucleation potential of amines has been shown to increase with basicity (Berndt et al., 2014; Kurtén et al., 2008; Yu et al., 2012). For example, although monopropylamine (MPA) and trimethylamine (TMA) are structural isomers of one another, MPA is likely to be a more potent nucleator due to its stronger basicity. The suite of alkylamines that have been commonly detected in the atmosphere contains multiple structural isomers (e.g. monoethylamine, MEA, and dimethylamine, DMA), making it difficult to speciate the amines using mass spectrometry (MS) without prior separation. Multiple field investigations sampling atmospheric particles using MS analysis have reported the detection of amine ion peaks but have been unable to assign them to a specific amine (Aiken et al., 2009; Denkenberger et al., 2007; Silva et al., 2008; Yao et al., 2016). Derivatization of alkylamines coupled with HPLC or GC separation has been reported to aid in separation and quantitation of amine species (Akyüz, 2007; Huang et al., 2009; Fournier et al., 2008; Key et al., 2011; Possanzini and Di Palo, 1990). However, these approaches are time consuming, require optimization of reaction conditions, and employ phase separations, which use large quantities of consumables, reagents, and solvents. Capillary electrophoresis has also been employed for aqueous amine separation, but in either case derivatization was required (Dabek-Zlotorzynska and Maruszak, 1998) or the

separation of atmospherically relevant cations was not addressed (Fekete et al., 2006). The use of ion chromatography to directly separate and quantify atmospheric alkylamines has been demonstrated (Chang et al., 2003; Dawson et al., 2014; Erupe et al., 2010; Huang et al., 2014; Li et al., 2009; Murphy et al., 2007; VandenBoer et al., 2012; Verrielle et al., 2012), yet the established IC methods struggle with coeluting cations (Huang et al., 2014; Murphy et al., 2007; VandenBoer et al., 2012; Verrielle et al., 2012) or they do not address a full suite of atmospherically relevant alkylamines and inorganic cations (Chang et al., 2003; Dawson et al., 2014; Erupe et al., 2010; Li et al., 2009).

In this work we demonstrate the separation and quantitation of the nine most abundant atmospheric alkylamines, two alkyl diamines, and six inorganic cations through the use of ion chromatography. We show (i) the separation method approach to maximizing peak resolution in the context of real-time atmospheric sampling and analysis; (ii) the effects of column temperature on amine coelution; (iii) the method precisions, accuracies, sensitivities, and limits of detection (LODs) for all alkylamine and inorganic cations; and (iv) application of the method to the complex matrix of atmospheric BB particle extracts to demonstrate method sensitivity and robustness.

2 Methods

2.1 Chemicals and materials

Inorganic cation stock solutions were prepared from a primary mixed cation standard concentrate (Dionex six-cation II, lot no. 150326, Thermo Scientific, Waltham, MA, USA) consisting of Li^+ , Na^+ , NH_4^+ , K^+ , Mg^{2+} , and Ca^{2+} chloride salts. Alkylamines (MMA (monomethylamine, 40 % w/w), DMA (40 % w/w), TMA (25 % w/w), MEA (70 % w/w), DEA (diethylamine, >99 % w/w), TEA (triethylamine, >99 % w/w), MPA (>99 % w/w), iMPA (isomono-propylamine, >99 % w/w) MBA (monobutylamine, 99.5 % w/w), MEtA (monoethanolamine, \geq 98 % w/w), DAB (1,4-diaminobutane, >98.5 % w/w), and DAP (1,5-diaminopentane, 95 % w/w)) were purchased from Sigma-Aldrich (Oakville, ON, Canada). Calibration standards were prepared by serial dilution in $>18.2 \text{ M}\Omega \times \text{cm}$ ultrapure deionized water (Barnstead Nanopure Infinity, Thermo Scientific, Waltham, MA, USA). Since these alkylamine species will largely be protonated in solution we will denote each of these species in their cationic form (i.e. NR_3H^+) henceforth when referring to the condensed phase.

2.2 Ion chromatography

A Thermo Scientific ICS-2100 Ion Chromatography System (Thermo Scientific, Mississauga, ON, Canada) utilizing Reagent-Free Ion Chromatography (RFICTM) components was used to develop the separation of the selected amines and

inorganic cations. A Thermo Scientific methanesulfonic acid (MSA) eluent generator cartridge (EGC III, P/N: 074535) was used in conjunction with an ultrapure deionized water reservoir to supply the eluent mobile phase with H_3O^+ ions as the competing exchanger. A continuously regenerated trap column (CR-CTC II, P/N: 066262) was attached in series to the eluent cartridge to remove cation contaminants from the eluent, thereby improving instrument detection limits. Samples were preconcentrated on a cation-exchange column (TCC-ULP1; $5 \times 23 \text{ mm}$, P/N: 063783) using a Thermo Scientific AS-DV autosampler to deliver the desired volume. Concentrated analytes were separated using Thermo Scientific CG19 ($4 \times 50 \text{ mm}$, P/N: 076027) and CS19 ($4 \times 250 \text{ mm}$, P/N: 076026) guard and analytical cation-exchange columns. The column effluent was passed through a suppressor operating in recycle mode (CERS 500, 4 mm) prior to detection of the analytes using a DS6 heated conductivity cell thermostated at 30°C . The eluent conductance was recorded at 5 Hz and the chromatographic peaks were analyzed using the ChromeleonTM 7 software package. A Thermo Scientific CG15 ($4 \times 50 \text{ mm}$, P/N: 052200) guard column was added inline later to attempt further improvement in analyte separation.

2.3 CS19 separation optimization

The gradient program used for the separation of methylamines, ethylamines, other alkylamines, and six inorganic cations on the CS19 cation-exchange column was optimized by combining analyte separation parameters from multiple isocratic elution runs at varying MSA concentrations (1–16 mM) and mobile-phase flow rates ($0.75\text{--}1.25 \text{ mL min}^{-1}$) at a column temperature of 30°C . Maximum peak resolution was optimized using an eluent gradient program and the column temperature was increased to resolve coeluting peaks (see Sect. 3.1.1).

Optimal separation of a suite of 15 cations was achieved using a mobile-phase flow rate of 1.25 mL min^{-1} and a column temperature of 55°C . The eluent gradient program is as follows: an initial MSA concentration of 1 mM held for 20 min, a step increase to 4 mM followed immediately by an exponential ramp to 10 mM over 10 min (Chromeleon curve factor of 7). The final concentration of 10 mM was held for an additional 5 min, yielding a total run time of 35 min. The IC was returned to initial conditions and re-equilibrated for 10 min as the next 1 mL sample aliquot was prepared for injection by the AS-DV. The suppressor current, optimized for this flow rate and the maximum eluent concentration in accordance with the calculation provided by the manufacturer, was set at 37 mA. The typical back pressure in the system at these conditions was 2100 psi. The Chromeleon method file for the method described above is detailed in Fig. S1 in the Supplement.

2.4 Quality assurance and quality control

Standards were prepared using Class A Corning polymethylpentene 50 (± 0.06) mL volumetric flasks that were rinsed four times with ethanol and eight times with ultrapure water prior to use. Standards were stored in 60 mL brown Nalgene polypropylene bottles that were pre-cleaned in a 10 % HCl bath, followed by eight sequential rinses with distilled and ultrapure water, respectively. The mixed amine standards and mixed inorganic cation standards were prepared separately and each cation standard set was composed of five calibration standards, two check standards, and an ultrapure deionized water blank. Ranges and related parameters are denoted by mass injected, as the preconcentration column negates the effect of volume. All amine calibration standards had a mass calibration range of 5–500 ng. The mass range for each inorganic cation calibration is as follows: Li^+ (0.82–16 ng), Na^+ (7.8–160 ng), NH_4^+ (8.4–170 ng), K^+ (26–520 ng), Mg^{2+} (6.4–130 ng), and Ca^{2+} (18–360 ng). All calibration curves contained trace inorganic cation impurities from the ultrapure deionized water source or holding vessels that fell below the lowest calibration standard and were corrected accordingly to allow for inter-day method performance comparison. Trace quantities of amines were not observed. An example of a calibration blank chromatogram is presented in Fig. S2.

Method precision for each methyl- and ethylamine cation was determined using standard calibration curves ($n = 9$) injected across 5 different days spanning 3 months. The precision for the propyl- and butylamines was determined using two standard calibration curves analyzed over 1 month. Method precision for Li^+ , Na^+ , NH_4^+ , and K^+ was assessed using calibrations ($n = 6$) from 3 separate days spanning 2 months. Precision for each cation was calculated using the standard deviation (σ) in the slope of the linear calibration curves. Check standards positioned between the two highest and the two lowest calibration standards for each cation were used to determine method accuracy across the calibration range. The low check standard was 15 times greater than the lowest standard and the high check standard was 150 times higher than the lowest standard. Accuracy was determined by the percent relative error between the known and calculated concentrations of the check standards. The LODs for the singly charged inorganic cations ($n = 4$) and methyl- and ethylamines ($n = 5$) were determined using calibration standard and calibration blank chromatograms from 3 or more separate days. The LODs for the propyl- and butylamines were determined using calibration standard and blank chromatograms from 2 separate days. The LODs are reported as concentrations resulting in a ratio of signal peak height to background noise of 3. The background noise was determined using the standard deviation of the conductance signal that fell within the retention time window for each analyte in their respective calibration blank chromatograms.

To assess the method robustness in the presence of a complex matrix the gradient method standard addition was performed on a subsample of a size-resolved BB particle extract (320–560 nm; see Sect. 2.5). Standard addition was performed by adding known quantities of methyl- and ethylamine solution to a 0.5 mL subsample of the extract followed by dilution to 5 mL. The amount of the methyl- and ethylamines added to the internal calibration matched that of the external calibration. The slope and retention times for the methyl- and ethylamines from the internal calibration were calculated and compared to those performed externally to quantify matrix effects. Discussion of the analytical performance of the CS19 gradient program is presented in Sect. 3.1.2.

2.5 Size-resolved BB sample analysis

A size-resolved particle sample from a BB plume was collected using a nanoMOUDI II (nano micro-orifice uniform-deposit impactor, model 122-R, MSP Corp., Shoreview, MN, USA) in St. John's, Newfoundland, on 6 July 2013. Satellite images of the plume smoke, HYSPLIT back trajectories, and measured $\text{PM}_{2.5}$ concentrations reported by Environment and Climate Change Canada indicate that these plumes originated from boreal forest fires in northern Québec and Labrador on 4 July 2013 and travelled via Labrador and the Gulf of St. Lawrence to the sampling site (Di Lorenzo and Young, 2016). The nanoMOUDI samples were collected on 13 aluminum substrate stages into size-resolved bins of atmospheric particles with a diameter range spanning 0.010–18 μm . Air was sampled continuously for 25.5 h at a flow rate of 30 L min^{-1} . A subsample of each aluminum substrate (10 % of the total substrate area) was extracted into a glass vial with 5 mL ultrapure deionized water by sonication (VWR Scientific Products/Aquasonic 150HT, Ultrasonic Water Bath) for 40 min. The extracts were filtered using a 0.2 μm polytetrafluoroethylene (PTFE) filter and stored in polypropylene vials at 4 °C prior to analysis by IC within 24 h. Cation analytes within these samples all fell within their respective calibration ranges and did not require any further dilution. An aluminum substrate field blank was also transported and exposed to the ambient atmosphere briefly at the collection site before being stored in a sealed container for the duration of the sample collection, transported back with the samples, and extracted simultaneously following the same procedure. All calculated quantities were corrected with measurements of the field blank with the additional error from this correction propagated into our final reported values. The field blank chromatogram for the size-resolved BB samples is presented in Fig. S2.

2.6 BC fire sample analysis

The full method for the collection and extraction of BB particle samples collected during July wildfires in British

Columbia is detailed in Di Lorenzo et al. (2017). Briefly, PM_{2.5} samples were collected at two sites approximately 100 km east of the BB location. The first site was located in Burnaby/Kensington Park (BKP) and the second was in North Vancouver/Second Narrows (NVSN). The particle samples were collected using beta attenuation particle monitors (5030 SHARP monitor at the BKP site, 5030i SHARP monitor at the NVSN site, Thermo Fisher Scientific, Waltham, MA, USA) at a 16.67 L min⁻¹ flow rate in 8 h intervals. Particles were collected on glass microfiber filter tape and stored at -20 °C until extracted. Approximately 37 % of each filtered particulate deposit was placed into a polypropylene vial with 10 mL of deionized water and sonicated for 40 min. The extracts were filtered with PTFE syringe filters (3 mm diameter, 0.2 µm pore size, VWR International, Radnor, Pennsylvania, USA) and diluted by a factor of 5 with ultrapure deionized water so that all analytes were in the IC calibration range. An unexposed area of the glass microfiber filter tape was sampled and extracted for use as a field blank. All calculated quantities and errors were blank-corrected using the field blank.

3 Results and discussion

3.1 Analytical method performance of CS19 cation-exchange column

3.1.1 Separation approach and optimization of parameters

Our approach to separation involved injecting the highest mixed inorganic cation and mixed amine standards for the expected working range (0.1–2.5 µg mL⁻¹) at static flow rates (0.75, 1, and 1.25 mL min⁻¹) while systematically increasing the isocratic eluent concentration (4–16 mM). The quality of each isocratic method was assessed by calculating the peak-to-peak resolution (R_s) using the retention time (t_r) and peak width at base (w) determined from the highest standard for each pair of cations following Eq. (1):

$$R_s = \frac{2(t_{r2} - t_{r1})}{w_2 + w_1}. \quad (1)$$

Using the upper limit of the expected working range for all analytes provides a lower limit on peak-to-peak resolution between these species. The peak-to-peak resolutions of the isocratic methods run using a 0.75 and 1.25 mL min⁻¹ flow rate for the selected inorganic and alkylamine cations are presented in Figs. S3 and S4. Peak-to-peak resolution between all peaks increased as the mobile-phase ionic strength was lowered when the flow rate was held constant. This is in agreement with Eq. (2), the fundamental resolution equation, which describes peak-to-peak resolution in terms of an efficiency factor (N), retention factor (k), and a selectivity factor

(α).

$$R = \left(\frac{\sqrt{N}}{4} \right) \left(\frac{k}{k+1} \right) \left(\frac{\alpha-1}{\alpha} \right). \quad (2)$$

With low mobile-phase ionic strength, the retention factor of the analytes is expected to increase, leading to greater resolution, consistent with our observations. In contrast, the effect of flow rate on peak resolution is nonintuitive and must be obtained empirically. Lower flow rates increase the retention factor, which in turn increases resolution. However, an increase in mobile-phase flow rate has a competing effect on the efficiency factor in Eq. (2). The efficiency term is governed by the theoretical plate height (H) as described by the Van Deemter equation (Eq. 3), which highlights the competing effect of flow rate (μ) on peak resolution:

$$H = A + \frac{B}{\mu} + C\mu. \quad (3)$$

Figures S3 and S4 show no loss in peak resolution when using a higher flow rate (1.25 mL min⁻¹ vs. 0.75 mL min⁻¹). To confirm that there was no loss in efficiency at higher flows, Van Deemter plots were created, using MMAH⁺ and TEAH⁺ as representative early and late-eluting species, by plotting theoretical plate height versus flow rate (Fig. S5). To do this, the theoretical plate heights described in Eq. (3) were calculated using Eq. (4), which relates H to column length (L), t_r and w . The A , B , and C terms of Eq. (3) were then determined by solving a system of equations using the calculated H values for MMAH⁺ and TEAH⁺ at three isocratic flow rates as they are located at opposite ends of the elution range of the six most abundant atmospheric alkylamines (Fig. 1).

$$\frac{L}{H} = 16 \left(\frac{t_r}{w} \right)^2 \quad (4)$$

The Van Deemter plots for MMAH⁺ and TEAH⁺ show no sacrifice in resolution when operating at higher flow rates and low eluent concentrations. The resolution between Mg²⁺ and Ca²⁺ as well as between TMAH⁺ and TEAH⁺ improved at the higher flow rate for all isocratic eluent concentrations. This was due to a decrease in peak width from diffusion band broadening. Furthermore, there was little to no sacrifice in resolution for all other cation peak pairs when operating at a higher flow rate. Of particular note, utilizing a 4 mM MSA isocratic separation at a 1.25 mL min⁻¹ flow rate instead of a 0.75 mL min⁻¹ flow resulted in a runtime that was 20 min shorter, which improves the applicability of IC for near-real-time analysis of hourly to bi-hourly atmospheric sample collection timescales. A shorter run time also improves the method throughput capacity for offline analyses and reduces total eluent consumption. For these reasons, the faster flow rate was selected in designing and optimizing a gradient program. Further isocratic methods using lower MSA concentrations (1 and 2 mM) were run at a

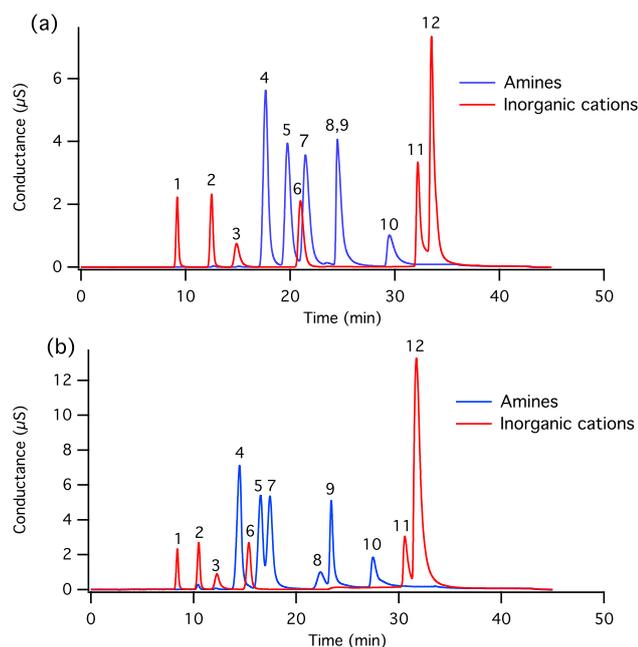


Figure 1. Separation of amine and inorganic cation standards with the highest-resolution gradient program at (a) 30 and (b) 55 °C. The order of elution and mass of cation injected in (a) is as follows: Li^+ (1, 16 ng), Na^+ (2, 158 ng), NH_4^+ (3, 169 ng), MMAH^+ (4, 500 ng), MEA^+ (5, 500 ng), K^+ (6, 524 ng), DMAH^+ (7, 500 ng), TMAH^+ (8, 500 ng), DEAH^+ (9, 500 ng), TEAH^+ (10, 500 ng), Mg^{2+} (11, 128 ng), and Ca^{2+} (12, 361 ng). Cation peaks represent the same mass injected and are labelled according to the same numeric identities in (b).

1.25 mL min^{-1} flow rate to quantify values of peak-to-peak resolution for the inorganic cations and alkylamines before approaching a gradient method (Fig. S3 and S4). An increase in resolution greater than one was observed for all analyte pairs aside from $\text{DEAH}^+/\text{TMAH}^+$ when using a 1 mM MSA eluent concentration.

All gradient methods that were tested started with a 1 mM hold, followed by a stepwise increase and/or ramp to higher eluent concentrations at a column temperature of 30 °C. By combining the best isocratic separations for various pairs of cation analytes sequentially, iterative modifications were used to improve resolution based on Eqs. (1)–(3) with the best separation method selected from amongst the iterations. Higher column temperature has been used to improve the quality of an IC separation method (Hatsis and Lucy, 2001; van Pinxteren et al., 2015; Rey and Pohl, 1996, 2003). Therefore, column temperature was systematically increased to investigate whether further improvement in peak-to-peak resolution was possible. Temperature effects on separation efficiency in IC are thermodynamically complex (Hatsis and Lucy, 2001; Kulis, 2004; Rey and Pohl, 1996), but they typically result in increased peak resolution because of improvements in mobile-phase diffusivity, which increases the ef-

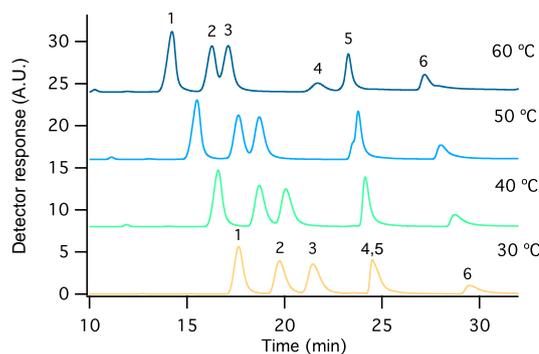


Figure 2. Separation of $1 \mu\text{g mL}^{-1}$ mixed amines standard with the final method gradient elution program at 30, 40, 50, and 60 °C. The peak elution order was MMAH^+ (1), MEA^+ (2), DMAH^+ (3), TMAH^+ (4), DEAH^+ (5), and TEAH^+ (6). The separation of diethylamine (DEA) from trimethylamine (TEA) was achieved at column temperatures greater than 50 °C.

ficiency term from Eq. (2). Higher temperatures can replicate the separation effects observed when adding an organic mobile-phase modifier (Hatsis and Lucy, 2001; Rey and Pohl, 1996). Figure 1a and b show gradient separations at 30 and 55 °C, respectively. At 30 °C, K^+ and DMAH^+ overlap considerably ($R_s = 0.45$) and DEAH^+ and TMAH^+ coelute. By increasing the column temperature to 55 °C, the extent of peak overlap between K^+ and neighbouring alkylamine cations is noticeably reduced ($R_s > 1$) and DEAH^+ and TMAH^+ are increasingly well resolved ($R_s = 1.48$). The effect of temperature on the separation of the alkylamines is demonstrated in Fig. 2, where the separation of DEAH^+ with TMAH^+ is achieved above 50 °C. The temperature increase also results in lower resolution between DMAH^+ and MEA^+ from $R_s = 1.57$ to $R_s = 1.08$. In our method, a column temperature of 55 °C produced peak-to-peak resolutions greater than a value of 1 between all six alkylamine cations and inorganic cations in the final gradient method, giving a 95 % separation between our target analytes and expected atmospheric interferences in the condensed phase. The peak-to-peak resolutions are summarized in Table 1. These represent a lower limit in peak resolution since they were calculated using peak parameters at the upper limit of the working range, which was determined based on maximum mixing ratios or mass loadings expected for the analysis of atmospheric samples containing these analytes.

The separation method produced in this work is able to overcome previously reported IC coelution difficulties between DEAH^+ and TMAH^+ and between MEA^+ and DMAH^+ (VandenBoer et al., 2012; Verrielle et al., 2012). Both DMA and TMA have been identified as dominant amines in emission studies, so it is important to achieve accurate and specific quantitation of both species in gas and particulate atmospheric samples (Facchini et al., 2008; Kuwata et al., 1983; Müller et al., 2009; Van Neste et al., 1987). Mul-

Table 1. Separation characteristics and statistics for the CS19 gradient method. The retention time (t_r) ranges for the methylamines, ethylamines, and inorganic cations were determined using retention time windows from a full calibration. The peak width and resolution were determined using the highest calibration standards for amines (500 ng) and inorganic (160–520 ng) cations. The t_r range and peak width were back-calculated for iMPAH⁺, MPAH⁺, MBAH⁺, DABH⁺, and DAPH⁺ based on the other alkylamine responses to column degradation. Sensitivity, precision, average LOD, and LOD range were analyzed using multiple calibration standards and blanks (see Sect. 2.4). Upper and lower range accuracies were assessed using high and low check standards for the alkylamines ($n = 6$) and inorganic cations ($n = 4$). The low check standards were 15 times more concentrated than the lowest calibration standard and the high check standards were 150 times more concentrated.

Cation	t_r (min)	Peak width (min)	Resolution	Sensitivity ($\mu\text{S min mol}^{-1}$)	Precision % ($\pm 1\sigma$)	Upper range accuracy (%)	Lower range accuracy (%)	Average LOD (pg; $\pm 1\sigma$)	LOD range (pg)
Li ⁺	8.2–8.4	0.72	2.68	11.5E08	2	96 \pm 5	82 \pm 4	0.6 \pm 0.2	0.3–0.8
Na ⁺	10.1–10.3	0.73	1.87	5.04E08	2	95 \pm 4	90 \pm 6	8 \pm 4	4–14
NH ₄ ⁺	11.8–12.1	1.18	0.65/1.85 ¹	2.45E08	4	103 \pm 4	50 \pm 50	22 \pm 17	7–47
MEtAH ⁺	12.7–13.0	0.99	0.56	2.2E08	–	–	–	3600 ($n = 1$)	–
MMAH ⁺	13.5–13.8	0.86	1.09	1.42E08	5	98 \pm 6	40 \pm 30	300 \pm 300	30–650
K ⁺	14.7–14.8	0.87	1.22	4.14E08	5	99 \pm 2	94 \pm 7	14 \pm 11	4–28
MEAH ⁺	15.5–15.8	0.79	1.08	0.90E08	7	97 \pm 5	40 \pm 10	500 \pm 200	200–700
DMAH ⁺	16.4–16.7	0.85	1.64	1.48E08	5	100 \pm 10	30 \pm 30	200 \pm 300	40–650
iMPAH ⁺	18.2–18.5	0.79	2.24	0.84E08	4	90 \pm 10	80 \pm 80	70 \pm 40	40–90
MPAH ⁺	20.1–20.4	0.88	1.55	0.62E08	12	88 \pm 4	90 \pm 90	50 \pm 40	20–80
TMAH ⁺	21.2–21.6	1.11	1.48	0.34E08	16	90 \pm 10	30 \pm 20	600 \pm 300	300–1200
DEAH ⁺	22.6–22.7	1.18	2.51	0.76E08	8	97 \pm 8	50 \pm 30	400 \pm 300	100–800
MBAH ⁺	25.3–25.6	0.43	3.12	0.62E08	1	80 \pm 20	100 \pm 80	910 \pm 30	890–930
TEAH ⁺	27.3–27.7	0.95	3.40	0.85E08	12	96 \pm 4	49 \pm 6	800 \pm 400	500–1400
Mg ²⁺	30.4–30.8	0.79	1.22	12.2E08	1	80 \pm 20	100 \pm 30	2000 \pm 3000	200–4000
Ca ²⁺	31.6–32.9	1.05	3.16	14.3E08	2	90 \pm 20	120 \pm 20	3700 \pm 200	3500–3800
DABH ⁺	36.6–36.9	1.48	0.98	4.5E08	–	–	–	1000 ($n = 1$)	–
DAPH ⁺	38.0–38.4	1.60	N/A	4.9E08	–	–	–	180 ($n = 1$)	–

¹ Resolution calculated for NH₄⁺ and MMAH⁺. – Insufficient replicates.

multiple field campaigns have detected large quantities of gas-phase and particle-phase MEA and DEA in ambient air as well (Facchini et al., 2008; Müller et al., 2009; Sorooshian et al., 2009; Yang et al., 2005, 2004). In some cases clean-up steps have been used to alleviate IC interferences from common atmospheric cation species in the quantitation of amines despite the fact that an 85 % evaporation loss of the amines, in addition to the extra sample handling, was reported when using a solid-phase extraction clean up (Huang et al., 2014). The CS19 IC method reported here is able to separate the most common atmospheric inorganic cations in addition to 11 common atmospheric amines (see separation of five additional alkylamines in Sect. 3.1.3). It can be easily applied to water-soluble atmospheric gas and particulate samples since they can be directly analyzed – without coelution or a clean-up step – with separation times of similar duration to many previously reported methods, including those employing an online IC method (Huang et al., 2014; Murphy et al., 2007; VandenBoer et al., 2012; Verrielle et al., 2012).

3.1.2 Instrumental performance and comparison for the methylamines, ethylamines, and inorganic cations

The performance statistics of the CS19 gradient method for each cation are summarized in Table 1. The method shows high reproducibility, with method precisions better

than 10 % for most analytes. Although the instrumental response varied from month-to-month for each analyte, this variability was random and the calibration curve slopes for each analyte showed no systematic decrease over time. The larger variability in the TMAH⁺ and TEAH⁺ calibration curves (± 16 and ± 11 %, respectively) is likely driven by their lower Henry's law constants (K_H) in water (Christie and Crisp, 1967), resulting in volatilization losses from standards. Concurrently, this variability could be driven by partitioning losses along the flow path, particularly when the tri-substituted amines reach the suppressor, which was not temperature controlled. In future investigations it may be worthwhile to acidify the standards to ensure the amines are maintained in their charged form in the aqueous phase. Alternatively, to combat losses to neutral forms, use of a salt converter suppressor accessory (Thermo Scientific, SC-CSRS 300, P/N: 067530), which keeps weak electrolytes in a separated sample fully protonated prior to their conductance measurement, may also aid in increasing long-term TMAH⁺ and TEAH⁺ precision.

The LODs for each analyte are reported in Table 1 as both a range and as the average LOD ($\pm 1\sigma$). The LODs are reported in this manner to reflect the high day-to-day variability in the calculated LODs. This variability may be driven by (i) the purity of the deionized water used for eluent generation; (ii) instrumental baseline noise and trace contamina-

tion on the day of analysis; and (iii) quality of labware cleaning prior to preparation of calibration blanks. Outliers in the LOD data set were found to result from trace contamination of analytical labware, sampling vials, or from systematic errors made in the preparation of standards or injection of samples on the IC (e.g. leaking autosampler caps, failing retention of concentrator column). The Grubb test was performed using a 95 % confidence interval to statistically identify outliers from LOD data sets. Calculated detection limits were determined to lie in the picogram per injection range for all analytes. The LODs for the inorganic cations were 10 to 100 times lower than those of the alkylamines, with the exception of Mg^{2+} and Ca^{2+} . Trace contamination of Ca^{2+} in our ultrapure deionized water led to higher LODs for the divalent cations. Our method shows high accuracy in the upper range of the calibrations for the methyl- and ethylamines, with accuracies ranging from 90 to 100 %. The accuracy was much lower for each methyl- and ethylamine cation at the low end of the calibration range where amine concentrations were approximately 1.5 times the limit of quantitation (LOQ). Quantitation near the method LOQ was more sensitive to small integration changes, which affect the calculated peak area, even when performing integrations manually, and this resulted in greater accuracy error. This is a drawback inherent to IC since wide analyte peaks are a result of persistently large stationary-phase particle sizes, causing band broadening via longer flow paths and increased diffusion during separation (i.e. the *A* and *B* terms in Eq. 3). The low alkylamine accuracies may also be driven by their air–water partitioning properties, which could result in losses during sample handling and during sample injection. The low and high range accuracies for all inorganic cations, with the exception of ammonium, were high (80–120 %) because concentrations were not near the LOQ for these analytes. The low check standard accuracy for ammonium likely arises due to similar issues as those discussed above for the tri-substituted amines.

To further test the efficacy of the separation method, a standard addition calibration was performed in the presence of the complex BB matrix. The calibration slopes and retention times for each analyte from the standard addition and external calibration performed on the same day are listed in Table S1. The slopes for the two calibrations varied between 0 and 8 %, which is within the method calibration precisions presented in Table 1. Thus, the BB sample extracts did not exhibit matrix effects. However, increasing retention times of approximately 0.3–0.5 min were observed for all cation analytes when performing the standard addition. This is an effect inherent in IC when samples with higher total quantities of cations are preconcentrated, resulting in a sample plug filling a greater quantity of the stationary-phase capacity. The initial weak mobile phase of the gradient method will therefore take a greater amount of time to elute all of the analyte cations from the preconcentration and analytical columns. This same increase in retention times is present in

the external calibration with increasingly concentrated standards (Table 1).

Previous IC instrumental precisions reported for use in quantifying the six atmospheric methyl- and ethylamines range from 0.4 to 17.2 %, which is comparable to our method (Table S2; Chang et al., 2003; Dawson et al., 2014; Erupe et al., 2010; Huang et al., 2014; Li et al., 2009; VandenBoer et al., 2012; Verrielle et al., 2012). Our separation method shows greater average variability than others due to our numerous assessments ($n = 9$) over multiple months, a more comprehensive analysis compared to previous reports. The sensitivity of this instrument is also similar to that of all other reported IC methods as the instrumental detection limits are in the picogram range. Only VandenBoer et al. (2012) and Chang et al. (2003) report lower detection limits and these are likely a result of a lower background signal from running the IC instruments online. Our method does not achieve instrumental limits of detection as low as those achieved using derivatization methods coupled with GC-MS or HPLC analysis (Akyüz, 2007; Fournier et al., 2008; Possanzini and Di Palo, 1990). However, multistep derivatization methods are prone to losses that must be quantified with internal standards. These losses can lead to higher overall method detection limits, which is not the case for direct analysis of water-soluble samples. Derivatization methods are also difficult to employ for near-real-time analyses of the atmosphere, making the approach less analytically attractive. Further, the IC separation method presented here is able to address additional matrix effects that may result from other atmospheric species through the use of a sample pre-concentration column. Only positively charged species are retained in this pre-concentration step and injected through the IC system for analysis, negating matrix effects from non-charged and anion species, as demonstrated by the standard addition to the BB sample extract.

Employing a method that is capable of quantifying amines at very low mixing ratios is valuable since recent work has shown that parts per quadrillion by volume (ppqv) concentrations of gaseous amines can lead to particle formation and growth (Almeida et al., 2013). If our method were applied to online atmospheric ambient sampling of gases or particles the method could be used to detect amines at ppqv mixing ratios. For example, a detectable signal for 100 ppqv mixing ratios could be attained by sampling through a bubbler, filter, or denuder at a low flow rate of 3 L min^{-1} for 1–10 h, depending on the amine. It may be possible to shorten the sampling duration to an hourly timescale if the method is interfaced with a high-sensitivity MS detector. Verrielle et al. (2014) observed a 5–30-fold improvement in method detection limits when interfacing their IC method with a MS detector.

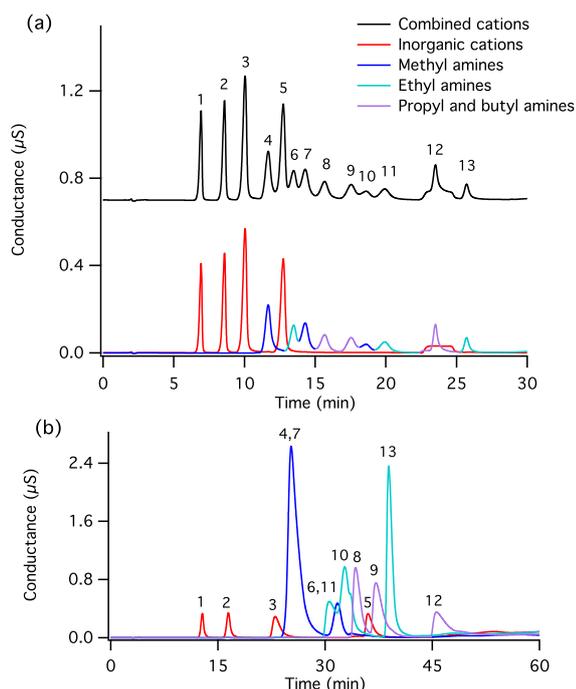


Figure 3. (a) Separation of amine and inorganic cation standards with the addition of MPAH⁺, iMPAH⁺, and MBAH⁺ using the final gradient program. The order of elution and mass of cation injected in (a) is as follows: Li⁺ (1, 1.6 ng), Na⁺ (2, 16 ng), NH₄⁺ (3, 17 ng), MMAH⁺ (4, 50 ng), K⁺ (5, 52 ng), MEAH⁺ (6, 50 ng), DMAH⁺ (7, 50 ng), iMPAH⁺ (8, 50 ng), MPAH⁺ (9, 50 ng), TMAH⁺ (10, 50 ng), DEAH⁺ (11, 50 ng), MBAH⁺ (12, 50 ng), and TEAH⁺ (13, 50 ng). (b) Separation of amine and inorganic cation standards with the addition of MPAH⁺, iMPAH⁺, and MBAH⁺ and the addition of the CG15 column using a modified gradient program. Cation peaks are labelled accordingly to the same identities in (b) and the mass of analyte injected is as follows: Li⁺ (1.6 ng), Na⁺ (16 ng), NH₄⁺ (17 ng), MMAH⁺ (500 ng), K⁺ (52 ng), MEAH⁺ (500 ng), DMAH⁺ (500 ng), iMPAH⁺ (500 ng), MPAH⁺ (500 ng), TMAH⁺ (500 ng), DEAH⁺ (500 ng), MBAH⁺ (500 ng), and TEAH⁺ (500 ng).

3.1.3 Expanded amine catalogue for other common atmospheric species

The separation method developed was further investigated to elucidate its utility in quantifying MPA, iMPA, and MBA, three amines that have been frequently detected in ambient air (Ge et al., 2011). In particular, this test was performed to assess their potential coelution with the fully separated methyl- and ethylamines. Without modification of the gradient method, we observed separation of these three amines from the original 12 cations with $R_s > 0.85$ (Fig. 3a). MPAH⁺ and iMPAH⁺ eluted between DMAH⁺ and TMAH⁺ and MBAH⁺ eluted later, but before TEAH⁺. The resolution is sufficient between all analyte peaks to allow quantitative analysis of the nine alkylamine cations and six inorganic cations. The separation statistics for these ad-

ditional amines are also presented in Table 1. Since the additional amines were injected after column degradation had occurred and retention times had noticeably shifted (see Fig. 3a vs. Fig. 1b and Sect. 3.1.5 for further discussion), retention time and peak width were estimated using changes in separation parameters from the original method development for the methyl- and ethylamines. Peak widths for the propyl- and butylamines were assumed to have increased by approximately 50 %, consistent for the same increases observed for the methyl- and ethylamines as a result of the column degradation. Retention times for MPAH⁺, iMPAH⁺, and MBAH⁺ were back-calculated to reflect the initial column conditions using these corrected peak widths and the resolution values determined from the chromatogram presented in Fig. 3. The method precisions for iMPAH⁺, MPAH⁺, and MBAH⁺ determined from two standard calibration injections ranged from 1 to 12 %. The reported average LODs for both propylamines are the lowest of the alkylamines, while MBAH⁺ has the highest method LOD because it elutes in a region with a high background due to the step change in eluent composition not being completely suppressed. The method accuracies for the three additional amines assessed by both the high and low check standards were within 80 % for all analytes. However, the large standard deviations in the accuracies for all low check standards highlights the challenge of method reproducibility for these analytes near the LOQ. This CS19 IC method can resolve three sets of alkylamine structural isomers, thus not only allowing full speciation of the suite of common atmospheric amines but also overcoming a limitation of direct MS analysis of the atmospheric matrix.

Since diamines have recently been shown to be potent sources of new particle formation and have been detected in field campaigns across the USA (Jen et al., 2016b), as well as near livestock, food processing factories, and sewage facilities (Ge et al., 2011), quantitation of DABH⁺ and DAPH⁺ was tested using the CS19 gradient method. The two diamines eluted after all 15 analytes, with DABH⁺ and Ca²⁺ having a resolution of 3.16 and DAPH⁺ and DABH⁺ having a resolution of 0.98. A single calibration curve was run to provide an estimate of the sensitivity and LOD for DABH⁺ and DAPH⁺ and the retention times for these two additional diamines were also back-calculated to reflect initial column conditions (Table 1). If the 10 mM MSA hold of the optimized method is extended by 5 min to give a total run time of 40 min, then DABH⁺ and DAPH⁺ can also be quantified. Ethanolamine (EAH⁺), a compound employed in industry and CO₂ capture (Ge et al., 2011), was also analyzed by this IC method and found to have poor resolution with NH₄⁺ and MMAH⁺ ($R_s < 0.65$, Table 1). However, the peak is still identifiable in samples containing these analytes.

3.1.4 Method development with the addition of an inline CG15 guard column

As mentioned previously, IC methods in the literature have been unable to separate potassium from the methyl- and ethylamines (Huang et al., 2014; VandenBoer et al., 2012), and in our current method K^+ has slight overlap with $MMAH^+$ ($R_s = 1.09$). We attempted to reduce peak overlap between K^+ and the alkylamines by adding a crown-ether-functionalized CG15 guard column inline after the CG19/CS19 columns. The addition of the CG15 column, which has increased selectivity for K^+ , resulted in an increased retention time for K^+ of 13 min. The best separation achieved using the additional guard column is shown in Fig. 3b, where K^+ still elutes within the alkylamine retention region. The gradient method used to achieve the separation used a flow rate of 1 mL min^{-1} , a column temperature of $55 \text{ }^\circ\text{C}$, and held a 1 mM MSA concentration for the first 30 min. The eluent concentration was step increased to 4 mM followed immediately by an exponential ramp to 10 mM over 20 min (Chromleon curve factor of 7). The final concentration of 10 mM is held for an additional 15 min, yielding a total run time of 65 min. Even when holding the initial MSA concentration at 1 mM for 50 min, the separation was unable to fully resolve the alkylamine peaks. An increase in retentivity for K^+ and NH_4^+ , as well as many of the alkylamines, indicated that the crown ether functionality was not selective for K^+ in this suite of analytes. With the addition of an organic modifier to the mobile phase or the ability to decrease column temperature, this increase in selectivity from the CG15 column might be harnessed to produce better separation. However, due to the limitations of the ICS-2100 system using RFIC we were unable to investigate these parameters. Furthermore, although a passable separation may be achieved when using a run time greater than 60 min, this would not be as applicable to online analyses as the CG/CS19 method developed without the addition of the CG15 column. A stationary phase similar to that of the CG/CS19 columns, but with some of this crown ether selectivity, could potentially yield better results than those presented here, particularly for the analysis of atmospheric samples containing large quantities of K^+ and amines.

3.1.5 Analytical column stability

Over the course of 5 months, peak retention times noticeably decreased and peak broadening of approximately 50 % occurred for all analytes. After more than 1000 sample and standard injections retention times had decreased by $1.9 \pm 0.1 \text{ min}$ depending on the cation. Peak-to-peak resolution, however, remained largely unchanged throughout the column degradation during standard and sample analysis, even with observed peak broadening. This is consistent with what has been previously reported in the literature when hundreds to thousands of injections have been run through an IC

column (VandenBoer et al., 2011). This may also be a result of column degradation from operating the CS19 column at a temperature higher than that recommended by the manufacturers.

During the course of method development severe peak broadening and subsequent peak-to-peak resolution loss of Mg^{2+} and Ca^{2+} was also observed. After the analysis of hundreds of samples the unresolved coeluting divalent cations had a peak width greater than 10 min wide. The Mg^{2+} and Ca^{2+} peak areas eventually became unresolved, with their combined peak area precision in the highest standard within 6 % ($\pm 1\sigma$) after 12 months of analysis. It was determined that this broadening effect observed for the Mg^{2+} and Ca^{2+} peaks was due to a malfunctioning suppressor. After replacing the suppressor, a peak-to-peak resolution greater than 1 was restored for these analytes. Furthermore, the cumulative analyte peak broadening that had occurred throughout method development and sample analysis for all the monovalent cations was also mitigated by installing the new suppressor. Retention times, however, were still shifted by $1.9 \pm 0.1 \text{ min}$, indicating that analytical column degradation had still occurred.

3.2 Biomass-burning particle analysis and discussion

3.2.1 Size-resolved alkylamines in particles from an aged biomass-burning plume

BB particles often contain a complex mixture of water-soluble ions, organics, elemental carbon, and other insoluble components, making them nonpareil for testing the robustness of an atmospheric measurement technique. Ions such as NH_4^+ and K^+ are consistently detected in BB plumes, regardless of sampling location as they are well characterized as being co-emitted species (Capes et al., 2008; Hudson et al., 2004; Pósfai et al., 2003). Particles released during forest fires have also been shown to contain highly oxidized large molecular weight organics (Di Lorenzo and Young, 2016; Saleh et al., 2014). We tested the robustness of our method on water-extracted aged BB particle samples collected by a cascade impactor in St. John's, Canada. Gas-sorption and reaction artifacts are minimized due to the gaseous flow path being directed around the nanoMOUDI impaction plates; therefore samples analyzed are representative of the particles in the atmosphere. An overlaid chromatogram of two different size-resolved particle samples (100–180 and 320–560 nm) shows the presence of $MMAH^+$, $DMAH^+$, and $DEAH^+$ in the aged BB samples with complete separation from K^+ , NH_4^+ , and Na^+ (Fig. 4). The maximum mass loadings for $MMAH^+$, $DMAH^+$, and $DEAH^+$ were found in particles with diameters of 320–560 nm and were 11 ± 3 , 208 ± 4 , and $1300 \pm 200 \text{ ng m}^{-3}$, respectively (Table S3). $TMAH^+$, MEA^+ , and $TEAH^+$ peaks were also detected in the BB size-resolved particle extracts. $TMAH^+$ and $TEAH^+$ reached mass loadings of $5 \pm 3 \text{ ng m}^{-2}$ and

$4 \pm 2 \text{ ng m}^{-3}$, respectively, while MEA^+ never exceeded a concentration of 1 ng m^{-3} in any size-resolved particle fraction (Table S3). The relative loadings of NH_4^+ , K^+ , or Na^+ to individual amines in these samples ranged from 2:1 to 1000:1, demonstrating the robustness of this method to quantify amines. Lobert et al. (1990) reported detecting C_1 – C_5 aliphatic amines from controlled BB experiments, which is consistent with our findings. BB-derived amines were also identified during the 2007 San Diego forest fires (Zauscher et al., 2013) with primary amines observed at approximately 6% by mass of organic content from an aged BB particle sample in British Columbia (Takahama et al., 2011). However, few studies have addressed the speciation and quantitation of alkylamines emitted from BB events. Schade and Crutzen (1995) estimated the emission rates for MMA, DMA, TMA, and MEA from BB sources based on controlled burn experiments, but they do not include a BB emission rate for DEA or TEA. These emission estimates are yet to include emission rates from atmospheric BB measurements (Lobert et al., 1990; Schade and Crutzen, 1995).

In Fig. 5a we show the molar ratio of the sum of the methyl- and ethylamines to ammonium which is considered to be the main atmospheric base, as a function of the size-resolved particles collected. The summed amine moles exceeded ammonium in the particle diameter range from 100 to 560 nm, and the ratio ranged from 0.50 to 1.9 in the fine mode ($\text{PM}_{1.0}$), with an average ratio of 0.92 in $\text{PM}_{1.0}$ calculated using nanoMOUDI bins up to this nominal cutoff. Quantities of NH_4^+ were below the detection limit in samples above $1 \mu\text{m}$, yielding no values for the ratio. The large error bars in the ratios are driven by the low molar quantities of ammonium in the samples as well as a higher variability in three extraction blanks (ultrapure deionized water sonicated in polypropylene tubes) injected on the day of analysis. For these reasons this method blank error was assigned to the size-resolved samples in place of the NH_4^+ error driven by the method precision and accuracy detailed in Table 1. To our knowledge, this is the first time that an amine-to-ammonium ratio greater than 1 has been reported in the $\text{PM}_{1.0}$ resolved fraction of atmospheric particles. An amine-to-ammonium ratio of 0.37 in fine-mode aerosol ($\text{PM}_{1.8}$) was reported by VandenBoer et al. (2011), but most reported ratios have been below 0.1 (Ge et al., 2011). The high ratios we observed can be attributed to large quantities of DEAH^+ and DMAH^+ . MMAH^+ was found to be in molar quantities 100 times less than that of ammonium while TMAH^+ , MEA^+ , and TEAH^+ were found to be in molar quantities 1000 to 10 000 times less than ammonium. Detecting such large molar ratio quantities of DEAH^+ and DMAH^+ to NH_4^+ in any particle sample is also unprecedented. Mono-substituted amines are the most frequently detected alkylamines in atmospheric particles and at molar ratios to ammonium of 1:100 or lower (Ge et al., 2011; Gorzelska et al., 1990; Mader et al., 2004; Müller et al., 2009; Yang et al., 2005, 2004; Zhang and Anastasio, 2003). In most instances where di-substituted or tri-

substituted amines have been identified in the particle phase, they are present at molar quantities equal to or less than the mono-substituted amines (Healy et al., 2015; Suzuki et al., 2001). Thus, such high quantities of DMAH^+ and DEAH^+ in these samples were unexpected and highly unusual compared to prior reports. In this case, the observation may be due to the fuel source of the fire or the interaction of the plume with a potent source of atmospheric amines. Previous work has identified di-substituted amines in large quantities from feedlot plumes (Sorooshian et al., 2008) and in marine particles (Facchini et al., 2008; Gibb et al., 1999a; Müller et al., 2009; Sorooshian et al., 2009; van Pinxteren et al., 2015; Youn et al., 2015). In fact, DMA and DEA have been reported as the second and third most abundant organic species in marine fine aerosol behind MSA during periods of high biological activity in the North Atlantic (Facchini et al., 2008). Other researchers have also suggested a moderate to high correlation between high biological activity and di-substituted amine particle mass loadings (Müller et al., 2009; Sorooshian et al., 2009). Laboratory investigations have shown that methylamines can be produced by marine phytoplankton degradation of quaternary amines to maintain an osmotic gradient as well as during periods of known zooplankton grazing (Gibb et al., 1999b). Based on the HYSPLIT back trajectories calculated for these samples (Di Lorenzo et al., 2017), it is possible that the BB plume particles interacted with gaseous DMA and DEA emitted from inland agricultural sources along the St. Lawrence River, coastal phytoplankton blooms, or with enhanced amine emissions in the coastal zone, which has been observed in the marine boundary layer of California (Youn et al., 2015). The high concentrations of DMA and DEA produced by marine biological activity could then partition into the BB particles and react to neutralize acids (e.g. sulfuric acid), form salts or amides with organic acids, or react with carbonyl moieties in the highly oxidized organic material produced via BB to form imines (Qiu and Zhang, 2013). If this explanation holds true then there may be preferential uptake of these amines over ammonia into the plume, as there is no evidence yet that suggests a larger agricultural or marine source of amines relative to ammonia to the atmosphere. A marine amine hypothesis, while consistent with observations in the literature, is beyond the scope of this work in terms of assigning a definitive DMA and DEA source.

3.2.2 Time series of amines in fresh biomass-burning plume particles from British Columbia

Our method was also applied to a time series of $\text{PM}_{2.5}$ samples collected at two different locations (BKP and NVSN) during a forest fire in the Lower Fraser Valley in British Columbia in the summer of 2015. These $\text{PM}_{2.5}$ samples were collected every 8 h while the plume was traversing each site and during collection the $\text{PM}_{2.5}$ concentration was in excess of $200 \mu\text{g m}^{-3}$. The relative ages for the smoke plumes sam-

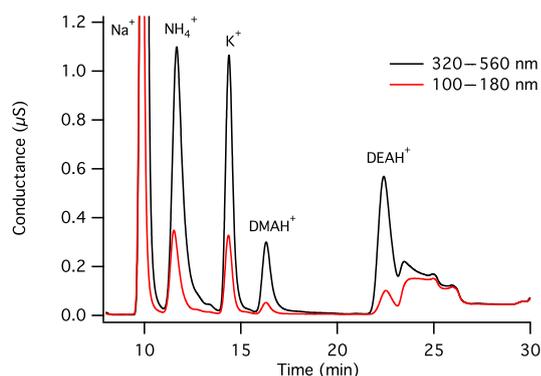


Figure 4. Overlaid chromatograms of MOUDI size-fractionated particle samples collected in St John's on 6 July 2013 during the intrusion of a biomass-burning plume that originated from northern Labrador and Québec. The robustness of the separation method for MMAH⁺, DMAH⁺, and DEAH⁺ from the common inorganic cations is demonstrated for the 320–560 (black) and 100–180 nm (red) size bins.

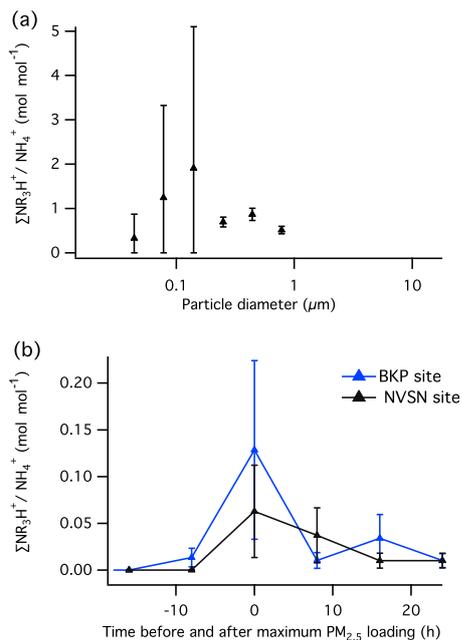


Figure 5. (a) Amine-to-ammonium ratio in the size-resolved aged biomass-burning sample originating from Québec and Labrador in the summer of 2013. (b) Amine-to-ammonium ratio for the Burnaby/Kensington Park (BKP) site and North Vancouver/Second Narrows (NVSN) site in British Columbia during the summer 2015 wildfires. The error bars in the graph represent propagated error in the amine and ammonium quantities resulting from variability in the field blanks and check standards during the analysis of the samples.

pled at both sites were calculated to be 20 h old, or less, and back trajectories indicated that the plume did not travel over the open ocean before being sampled (Di Lorenzo et al., 2017). In these test samples, the method was again able

to detect the presence of four different amines (iMPAH⁺, TMAH⁺, DEAH⁺, and TEAH⁺) with loadings of Na⁺, NH₄⁺, and K⁺ at ratios in excess of 100:1 (Fig. S6). Furthermore, the method was not only able to determine the presence of iMPAH⁺ but also differentiate it from MPAH⁺ and TMAH⁺, its two structural isomers. iMPAH⁺, TMAH⁺, DEAH⁺, and TEAH⁺ had maximum mass loadings in these fresher BB samples of 60 ± 40 , 9 ± 7 , 1.6 ± 0.8 , and $0.2 \pm 0.1 \text{ ng m}^{-3}$, respectively (Table S4). iMPAH⁺ was the amine detected in the largest molar quantities at both sampling sites in British Columbia. The detection of iMPAH⁺ has not previously been reported in BB particles, and based on our measurements may be important to quantify in future controlled burn experiments. Our results differ from the study conducted by Takahama et al. (2011) on the 2009 forest fires in British Columbia that reports the detection of primary amine groups, which further suggest that amine emissions from BB and/or their incorporation into BB particles are not well understood. Although our observed suite of amines includes iMPAH⁺, there was no indication of other primary amines from the analyses of the BB particles. The calculated quantities of amines for these samples may be biased high or low because the beta attenuation monitor used to sample the BB particles can be prone to gas-phase blow-on or blow-off artifacts via sorption or reactive mechanisms, as per traditional filter sampling strategies. Although the extracted field blank was corrected for any positive sorption biases on the filter tape (e.g. Müller et al., 2009), it was unable to account for any reactive uptake of gas-phase amines during the sampling period. Therefore, the filter tape BB samples presented here represent a time-integrated particulate composition assuming thermodynamic equilibrium between the gas and particulate phases for the duration of collection for each sample.

A time series of the amine-to-ammonium molar ratio as the smoke plume intrudes into both the BKP and NVSN sites is presented in Fig. 5b. There were either no amines present or they were present in concentrations below our detection limits in the ambient particles collected on the front edge of the plume intrusions. When the maximum PM_{2.5} mass loading of the plume reached the sampling site at $t = 0$, we saw an absolute maxima in total amine concentration as well as a relative maxima in the particulate amine-to-ammonium molar ratio (Fig. 5b). The particulate amine concentrations and the amine-to-ammonium ratio then tapered off as the plume diluted and passed through the site. The measured amine-to-ammonium ratio in these samples is consistent with previously reported literature values (Ge et al., 2011). The measured amine species and quantities, as in the aged plume, could be indicative of the BB source fuel, fire type, or amine levels in air masses intercepted that were subsequently incorporated by partitioning and reacting into the condensed phase. Since the smoke plumes were less than a day old and the plume did not travel over the open ocean, it is less likely that offshore marine amine emissions interacted with the plume. However, the BKP and NVSN sampling sites

are positioned in a coastal urban centre and anthropogenic amine emissions from industry or animal husbandry operations nearby, as well as coastal amine emissions, may still have been incorporated into the plume before it was sampled.

4 Conclusions

We developed an ion chromatographic method that can separate and quantify the nine most abundant atmospheric alkylamines and two alkyl diamines from common inorganic atmospheric cations. Ion chromatography methods reported in the literature cannot fully resolve alkylamine peaks nor separate interferences from K^+ , Mg^{2+} , and Ca^{2+} . In this work, we report the ability to overcome these prevalent issues for atmospheric sampling with a rapidity that can also be applied to near-real-time analyses of aqueous atmospheric extracts by IC. Additionally, the method is able to separate and quantify three sets of structural isomers, a limitation for direct particle and gas sampling mass spectrometry instrumentation in quantifying atmospheric alkylamines. The method detection limits are comparable to other published IC methods in the literature, but the described method is not as sensitive as instrumentation used in conjunction with derivatization methods coupled with GC-MS or LC-MS, which can suffer from sample processing losses.

The IC method is robust. Two sets of BB particle samples collected at two different locations in Canada were injected onto the IC column and the method detected and quantified amines in the presence of a complex matrix where inorganic analytes, such as K^+ , reached ratios of 1000:1 relative to the alkylamines. A standard addition demonstrated that the BB matrix does not have any influence on the ability of the IC method to quantify these analytes. This is a major improvement over all prior reports of the application of IC to the detection of amines in aqueous extracts of atmospheric particulate matter. Our results suggest that increasing focus on speciation and quantitation of various alkylamines from direct BB emissions and their propensity to undergo reactive uptake with BB particles is needed to constrain global budgets of atmospheric sources and fate of alkylamines.

Overall, the developed IC method shows promise for (i) adoption into standard analysis of water-soluble atmospheric extracts; (ii) incorporation into online instrumentation already using ion chromatography for near real-time analysis of water-soluble atmospheric samples; and (iii) interfacing with mass spectrometry for even higher analytical sensitivity, particularly where supporting measurements for ppqv levels of amines may be stimulating new particle formation in the atmosphere.

Data availability. To access the data sets presented and discussed, please see the Supplement or contact the corresponding author.

The Supplement related to this article is available online at doi:10.5194/amt-10-1061-2017-supplement.

Author contributions. Trevor C. VandenBoer designed the experiments and Bryan K. Place, Aleya T. Quilty, and Robert A. Di Lorenzo carried them out. Bryan K. Place prepared the manuscript with contributions from all coauthors.

Competing interests. The authors declare that they have no conflict of interest.

Acknowledgements. The authors thank Geoff Doerksen at the Lower Fraser Valley Air Quality Monitoring Network in British Columbia for supplying biomass burning samples. Thanks to Joseph Bautista for help collecting biomass burning samples in St. John's, as well as Jamie Warren and Kathryn Dawe for their assistance in method development. Finally, the authors would like to thank Cora Young for her helpful comments in the writing of the manuscript. TCV was supported by a Government of Canada Banting Postdoctoral Fellowship and the ICS-2100 was procured through the Canadian Foundation for Innovation. Funding for this work was provided by the Government of Newfoundland and Labrador Department of Forestry and Agrifoods through a Centre for Forestry Science and Innovation grant (project no. 221269).

Edited by: P. Herckes

Reviewed by: three anonymous referees

References

- Aiken, A. C., Salcedo, D., Cubison, M. J., Huffman, J. A., DeCarlo, P. F., Ulbrich, I. M., Docherty, K. S., Sueper, D., Kimmel, J. R., Worsnop, D. R., Trimborn, A., Northway, M., Stone, E. A., Schauer, J. J., Volkamer, R. M., Fortner, E., de Foy, B., Wang, J., Laskin, A., Shutthanandan, V., Zheng, J., Zhang, R., Gaffney, J., Marley, N. A., Paredes-Miranda, G., Arnott, W. P., Molina, L. T., Sosa, G., and Jimenez, J. L.: Mexico City aerosol analysis during MILAGRO using high resolution aerosol mass spectrometry at the urban supersite (T0) – Part 1: Fine particle composition and organic source apportionment, *Atmos. Chem. Phys.*, 9, 6633–6653, doi:10.5194/acp-9-6633-2009, 2009.
- Akyüz, M.: Simultaneous determination of aliphatic and aromatic amines in indoor and outdoor air samples by gas chromatography-mass spectrometry, *Talanta*, 71, 486–492, doi:10.1016/j.talanta.2006.10.028, 2007.
- Almeida, J., Schobesberger, S., Kürten, A., Ortega, I. K., Kupiainen-Määttä, O., Praplan, A. P., Adamov, A., Amorim, A., Bianchi, F., Breitenlechner, M., David, A., Dommen, J., Donahue, N. M., Downard, A., Dunne, E., Duplissy, J., Ehrhart, S., Flagan, R. C., Franchin, A., Guida, R., Hakala, J., Hansel, A., Heinritzi, M., Henschel, H., Jokinen, T., Junninen, H., Kajos, M., Kangasluoma, J., Keskinen, H., Kupc, A., Kurtén, T., Kvashin, A. N., Laaksonen, A., Lehtipalo, K., Leiminger, M., Leppä, J., Loukonen, V., Makhmutov, V., Mathot, S., McGrath, M. J.,

- Nieminen, T., Olenius, T., Onnela, A., Petäjä, T., Riccobono, F., Riipinen, I., Rissanen, M., Rondo, L., Ruuskanen, T., Santos, F. D., Sarnela, N., Schallhart, S., Schnitzhofer, R., Seinfeld, J. H., Simon, M., Sipilä, M., Stozhkov, Y., Stratmann, F., Tomé, A., Tröstl, J., Tsagkogeorgas, G., Vaattovaara, P., Viisanen, Y., Virtanen, A., Vrtala, A., Wagner, P. E., Weingartner, E., Wex, H., Williamson, C., Wimmer, D., Ye, P., Yli-Juuti, T., Carslaw, K. S., Kulmala, M., Curtius, J., Baltensperger, U., Worsnop, D. R., Vehkamäki, H., and Kirkby, J.: Molecular understanding of sulphuric acid-amine particle nucleation in the atmosphere, *Nature*, 502, 359–363, doi:10.1038/nature12663, 2013.
- Angelino, S., Suess, D. T., and Prather, K. A.: Formation of aerosol particles from reactions of secondary and tertiary alkylamines: Characterization by aerosol time-of-flight mass spectrometry, *Environ. Sci. Technol.*, 35, 3130–3138, doi:10.1021/es0015444, 2001.
- Ault, A. P., Moffet, R. C., Baltrusaitis, J., Collins, D. B., Ruppel, M. J., Cuadra-Rodriguez, L. A., Zhao, D., Guasco, T. L., Ebben, C. J., Geiger, F. M., Bertram, T. H., Prather, K. A., and Grassian, V. H.: Size-dependent changes in sea spray aerosol composition and properties with different seawater conditions, *Environ. Sci. Technol.*, 47, 5603–5612, doi:10.1021/es400416g, 2013.
- Barsanti, K. C., McMurry, P. H., and Smith, J. N.: The potential contribution of organic salts to new particle growth, *Atmos. Chem. Phys.*, 9, 2949–2957, doi:10.5194/acp-9-2949-2009, 2009.
- Berndt, T., Stratmann, F., Sipilä, M., Vanhanen, J., Petäjä, T., Mikkilä, J., Grüner, A., Spindler, G., Lee Mauldin III, R., Curtius, J., Kulmala, M., and Heintzenberg, J.: Laboratory study on new particle formation from the reaction OH + SO₂: influence of experimental conditions, H₂O vapour, NH₃ and the amine tert-butylamine on the overall process, *Atmos. Chem. Phys.*, 10, 7101–7116, doi:10.5194/acp-10-7101-2010, 2010.
- Berndt, T., Sipilä, M., Stratmann, F., Petäjä, T., Vanhanen, J., Mikkilä, J., Patokoski, J., Taipale, R., Mauldin III, R. L., and Kulmala, M.: Enhancement of atmospheric H₂SO₄ / H₂O nucleation: organic oxidation products versus amines, *Atmos. Chem. Phys.*, 14, 751–764, doi:10.5194/acp-14-751-2014, 2014.
- Boucher, O., Randall, D., Artaxo, P., Bretherton, C., Feingold, G., Forster, P., Kerminen, V.-M., Kondo, Y., Liao, H., Lohmann, U., Rasch, P., Sathesh, S. K., Sherwood, S., Stevens, B. and Zhang, X. Y.: Clouds and Aerosols, in: *Climate Change 2013: The Physical Science Basis, Contribution of Working Group I to the Fifth Assessment Report of the Intergovernmental Panel on Climate Change*, Cambridge, United Kingdom and New York, NY, USA, 87 pp., 2013.
- Bzdek, B. R., Ridge, D. P., and Johnston, M. V.: Amine exchange into ammonium bisulfate and ammonium nitrate nuclei, *Atmos. Chem. Phys.*, 10, 3495–3503, doi:10.5194/acp-10-3495-2010, 2010.
- Bzdek, B. R., Ridge, D. P., and Johnston, M. V.: Amine reactivity with charged sulfuric acid clusters, *Atmos. Chem. Phys.*, 11, 8735–8743, doi:10.5194/acp-11-8735-2011, 2011.
- Cadle, S. H. and Mulawa, P. A.: Low-molecular-weight aliphatic amines in exhaust from catalyst-equipped cars, *Environ. Sci. Technol.*, 14, 718–723, doi:10.1021/es60166a011, 1980.
- Capes, G., Johnson, B., McFiggans, G., Williams, P. I., Haywood, J., and Coe, H.: Aging of biomass burning aerosols over West Africa: Aircraft measurements of chemical composition, microphysical properties, and emission ratios, *J. Geophys. Res.-Atmos.*, 113, 1–13, doi:10.1029/2008JD009845, 2008.
- Christie, A. O. and Crisp, D. J.: Activity coefficients of the n-primary, secondary and tertiary aliphatic amines in aqueous solution, *J. Appl. Chem.*, 17, 11–14, 1967.
- Chang, I.-H., Lee, C.-G., and Lee, D. S.: Development of an automated method for simultaneous determination of low molecular weight aliphatic amines and ammonia in ambient air by diffusion scrubber coupled to ion chromatography, *Anal. Chem.*, 75, 6141–6146, doi:10.1021/ac0347314, 2003.
- Creamean, J. M., Ault, A. P., Ten Hoeve, J. E., Jacobson, M. Z., Roberts, G. C., Prather, K. A., Hoeve, J. E. Ten, Jacobson, Z., Roberts, G. C., and Prather, K. A.: Measurements of aerosol chemistry during new particle formation events at a remote rural mountain site, *Environ. Sci. Technol.*, 45, 8208–8216, doi:10.1021/es103692f, 2011.
- Dabek-Zlotorzynska, E. and Maruszak, W.: Determination of dimethylamine and other low-molecular-mass amines using capillary electrophoresis with laser-induced fluorescence detection, *J. Chromatogr. B*, 714, 77–85, 1998.
- Dall’Osto, M., Ceburnis, D., Monahan, C., Worsnop, D. R., Bialek, J., Kulmala, M., Kurtén, T., Ehn, M., Wenger, J., Sodeau, J., Healy, R., and O’Dowd, C.: Nitrogenated and aliphatic organic vapors as possible drivers for marine secondary organic aerosol growth, *J. Geophys. Res.-Atmos.*, 117, 1–10, doi:10.1029/2012JD017522, 2012.
- Dawson, M. L., Perraud, V., Gomez, A., Arquero, K. D., Ezell, M. J., and Finlayson-Pitts, B. J.: Measurement of gas-phase ammonia and amines in air by collection onto an ion exchange resin and analysis by ion chromatography, *Atmos. Meas. Tech.*, 7, 2733–2744, doi:10.5194/amt-7-2733-2014, 2014.
- Denkenberger, K. A., Moffet, R. C., Holecek, J. C., Rebotier, T. P., and Prather, K. A.: Real-time, single-particle measurements of oligomers in aged ambient aerosol particles, *Environ. Sci. Technol.*, 41, 5439–5446, doi:10.1021/es0703291, 2007.
- Di Lorenzo, R. A. and Young, C. J.: Size separation method for absorption characterization in brown carbon: Application to an aged biomass burning sample, *Geophys. Res. Lett.*, 43, 458–465, doi:10.1002/2015GL066954, 2016.
- Di Lorenzo, R. A., Washenfelder, R. A., Attwood, A. R., Guo, H., Xu, L., Ng, N. L., Weber, R. J., Baumann, K., Edgerton, E., and Young, C. J.: Molecular size separated brown carbon absorption for fresh and aged biomass burning plumes at multiple field sites, *Environ. Sci. Technol.*, in press, doi:10.1021/acs.est.6b06160, 2017.
- Ehn, M., Thornton, J. A., Kleist, E., Sipilä, M., Junninen, H., Pullinen, I., Springer, M., Rubach, F., Tillmann, R., Lee, B., Lopez-Hilfiker, F., Andres, S., Acir, I.-H., Rissanen, M., Jokinen, T., Schobesberger, S., Kangasluoma, J., Kontkanen, J., Nieminen, T., Kurtén, T., Nielsen, L. B., Jørgensen, S., Kjaergaard, H. G., Canagaratna, M., Maso, M. D., Berndt, T., Petäjä, T., Wahner, A., Kerminen, V.-M., Kulmala, M., Worsnop, D. R., Wildt, J., and Mentel, T. F.: A large source of low-volatility secondary organic aerosol, *Nature*, 506, 476–479, doi:10.1038/nature13032, 2014.
- Erupe, M. E., Liberman-Martin, A., Silva, P. J., Malloy, Q. G. J., Yonis, N., Cocker, D. R., and Purvis-Roberts, K. L.: Determination of methylamines and trimethylamine-N-oxide in particulate

- matter by non-suppressed ion chromatography, *J. Chromatogr. A*, 1217, 2070–2073, doi:10.1016/j.chroma.2010.01.066, 2010.
- Erupe, M. E., Viggiano, A. A., and Lee, S.-H.: The effect of trimethylamine on atmospheric nucleation involving H₂SO₄, *Atmos. Chem. Phys.*, 11, 4767–4775, doi:10.5194/acp-11-4767-2011, 2011.
- Facchini, M. C., Decesari, S., Rinaldi, M., Carbone, C., Finessi, E., Mircea, M., Fuzzi, S., Moretti, F., Tagliavini, E., Ceburnis, D., and O'Dowd, C. D.: Important source of marine secondary organic aerosol from biogenic amines, *Environ. Sci. Technol.*, 42, 9116–9121, doi:10.1021/es8018385, 2008.
- Fekete, A., Frommberger, M., Ping, G., Lahaniatis, M. R., Lintelman, J., Fekete, J., Gebefugi, I., Malik, A. K., Kettrup, A., and Schmitt-Kopplin, P.: Development of a capillary electrophoretic method for the analysis of low-molecular-weight amines from metal working fluid aerosols and ambient air, *Electrophoresis*, 27, 1237–1247, doi:10.1002/elps.200500724, 2006.
- Fournier, M., Lesage, J., Ostiguy, C., and Tra, H. V.: Sampling and analytical methodology development for the determination of primary and secondary low molecular weight amines in ambient air, *J. Env. Monit.*, 10, 379–386, doi:10.1039/b719091n, 2008.
- Ge, X., Wexler, A. S., and Clegg, S. L.: Atmospheric amines – Part I. A review, *Atmos. Environ.*, 45, 524–546, doi:10.1016/j.atmosenv.2010.10.012, 2011.
- Gibb, S. W., Mantoura, R. F. C., Liss, P. S., and Barlow, R. G.: Distributions and biogeochemistries of methylamines and ammonium in the Arabian Sea, *Deep-Sea Res. Pt. II*, 46, 593–615, doi:10.1016/S0967-0645(98)00119-2, 1999a.
- Gibb, S. W., Mantoura, R. F. C., and Liss, P. S.: Ocean-atmosphere exchange and atmospheric speciation of ammonia and methylamines in the region of the NW Arabian Sea, *Global Biogeochem. Cy.*, 13, 161–178, doi:10.1029/98GB00743, 1999b.
- Gorzelska, K. and Galloway, J. N.: Amine nitrogen in the atmospheric environment over the North Atlantic Ocean, *Global Biogeochem. Cy.*, 4, 309–333, doi:10.1029/GB004i003p00309, 1990.
- Hatsis, P. and Lucy, C. A.: Evaluation of column temperature as a means to alter selectivity in the cation exchange separation of alkali metals, alkaline earth metals and amines, *Analyst*, 126, 2113–2118, doi:10.1039/b106639k, 2001.
- Healy, R. M., Evans, G. J., Murphy, M., Sierau, B., Arndt, J., McGillicuddy, E., O'Connor, I. P., Sodeau, J. R., and Wenger, J. C.: Single-particle speciation of alkylamines in ambient aerosol at five European sites, *Anal. Bioanal. Chem.*, 407, 5899–5909, doi:10.1007/s00216-014-8092-1, 2015.
- Hodshire, A. L., Lawler, M. J., Zhao, J., Ortega, J., Jen, C., Yli-Juuti, T., Brewer, J. F., Kodros, J. K., Barsanti, K. C., Hanson, D. R., McMurry, P. H., Smith, J. N., and Pierce, J. R.: Multiple new-particle growth pathways observed at the US DOE Southern Great Plains field site, *Atmos. Chem. Phys.*, 16, 9321–9348, doi:10.5194/acp-16-9321-2016, 2016.
- Huang, G., Hou, J., and Zhou, X.: A measurement method for atmospheric ammonia and primary amines based on aqueous sampling, OPA derivatization and HPLC analysis, *Environ. Sci. Technol.*, 43, 5851–5856, doi:10.1021/es900988q, 2009.
- Huang, R.-J., Li, W.-B., Wang, Y.-R., Wang, Q. Y., Jia, W. T., Ho, K.-F., Cao, J. J., Wang, G. H., Chen, X., El Haddad, I., Zhuang, Z. X., Wang, X. R., Prévôt, A. S. H., O'Dowd, C. D., and Hoffmann, T.: Determination of alkylamines in atmospheric aerosol particles: a comparison of gas chromatography-mass spectrometry and ion chromatography approaches, *Atmos. Meas. Tech.*, 7, 2027–2035, doi:10.5194/amt-7-2027-2014, 2014.
- Hudson, P. K., Murphy, D. M., Cziczko, D. J., Thomson, D. S., de Gouw, J. A., Warneke, C., Holloway, J., Jost, H. J., and Hübler, G.: Biomass-burning particle measurements: Characteristics composition and chemical processing, *J. Geophys. Res.-Atmos.*, 109, 1–11, doi:10.1029/2003JD004398, 2004.
- IARC: Monographs on the Evaluation of Carcinogenic Risks to Humans: Outdoor Air Pollution (World Health Organization), available at: <http://monographs.iarc.fr/ENG/Classification>, last access: 9 August 2016.
- Jen, C. N., Zhao, J., McMurry, P. H., and Hanson, D. R.: Chemical ionization of clusters formed from sulfuric acid and dimethylamine or diamines, *Atmos. Chem. Phys.*, 16, 12513–12529, doi:10.5194/acp-16-12513-2016, 2016a.
- Jen, C. N., Bachman, R., Zhao, J., McMurry, P. H., and Hanson, D. R.: Diamine-sulfuric acid reactions are a potent source of new particle formation, *Geophys. Res. Lett.*, 43, 867–873, doi:10.1002/2015GL066958, 2016b.
- Key, D., Stihle, J., Petit, J.-E., Bonnet, C., Depernon, L., Liu, O., Kennedy, S., Latimer, R., Burgoyne, M., Wanger, D., Webster, A., Casunuran, S., Hidalgo, S., Thomas, M., Moss, J. A., and Baum, M. M.: Integrated method for the measurement of trace nitrogenous atmospheric bases, *Atmos. Meas. Tech.*, 4, 2795–2807, doi:10.5194/amt-4-2795-2011, 2011.
- Kovač, N., Glavaš, N., Dolenec, M., Šmuc, N. R., and Šlejkovec, Z.: Chemical Composition of Natural Sea Salt from the Secovlje Salina (Gulf of Trieste, northern Adriatic), *Acta Chim. Slov.*, 60, 706–714, 2013.
- Kuhn, U., Sintermann, J., Spirig, C., Jocher, M., Ammann, C., and Neftel, A.: Basic biogenic aerosol precursors: Agricultural source attribution of volatile amines revised, *Geophys. Res. Lett.*, 38, 1–8, doi:10.1029/2011GL047958, 2011.
- Kulisa, K.: The effect of temperature on the cation-exchange separations in ion chromatography and the mechanism of zone spreading, *Chem. Anal.*, 49, 665–689, 2004.
- Kulmala, M., Kontkanen, J., Junninen, H., Lehtipalo, K., Manninen, H. E., Nieminen, T., Petäjä, T., Sipilä, M., Schobesberger, S., Rantala, P., Franchin, A., Jokinen, T., Järvinen, E., Äijälä, M., Kangasluoma, J., Hakala, J., Aalto, P. P., Paasonen, P., Mikkilä, J., Vanhanen, J., Aalto, J., Hakola, H., Makkonen, U., Ruuskanen, T., Mauldin, R. L., Duplissy, J., Vehkamäki, H., Bäck, J., Kortelainen, A., Riipinen, I., Kurtén, T., Johnston, M. V., Smith, J. N., Ehn, M., Mentel, T. F., Lehtinen, K. E. J., Laaksonen, A., Kerminen, V.-M., and Worsnop, D. R.: Direct observations of atmospheric aerosol nucleation, *Science*, 339, 943–946, doi:10.1126/science.1227385, 2013.
- Kurtén, T., Loukonen, V., Vehkamäki, H., and Kulmala, M.: Amines are likely to enhance neutral and ion-induced sulfuric acid-water nucleation in the atmosphere more effectively than ammonia, *Atmos. Chem. Phys.*, 8, 4095–4103, doi:10.5194/acp-8-4095-2008, 2008.
- Kürten, A., Bergen, A., Heinritzi, M., Leiminger, M., Lorenz, V., Piel, F., Simon, M., Sitals, R., Wagner, A. C., and Curtius, J.: Observation of new particle formation and measurement of sulfuric acid, ammonia, amines and highly oxidized organic molecules at a rural site in central Germany, *Atmos. Chem. Phys.*, 16, 12793–12813, doi:10.5194/acp-16-12793-2016, 2016.

- Kuwata, K., Akiyama, E., Yamasaki, Y., Yamasaki, H., Kuge, Y., and Kiso, Y.: Trace determination of low molecular weight aliphatic amines in air by gas chromatography, *Anal. Chem.*, 55, 2199–2201, doi:10.1021/ac00263a046, 1983.
- Leach, J., Blanch, A., and Bianchi, A. C.: Volatile organic compounds in an urban airborne environment adjacent to a municipal incinerator, waste collection centre and sewage treatment plant, *Atmos. Environ.*, 33, 4309–4325, doi:10.1016/S1352-2310(99)00115-6, 1999.
- Li, F., Liu, H., Xue, C., Xin, X., Xu, J., Chang, Y., Xue, Y., and Yin, L.: Simultaneous determination of dimethylamine, trimethylamine and trimethylamine-n-oxide in aquatic products extracts by ion chromatography with non-suppressed conductivity detection, *J. Chromatogr. A*, 1216, 5924–5926, doi:10.1016/j.chroma.2009.06.038, 2009.
- Lloyd, J. A., Heaton, K. J., and Johnston, M. V.: Reactive uptake of trimethylamine into ammonium nitrate particles, *J. Phys. Chem. A*, 113, 4840–4843, doi:10.1021/jp900634d, 2009.
- Lobert, J. M., Scharffe, D. H., Hao, W. M., and Crutzen, P. J.: Importance of biomass burning in the atmospheric budgets of nitrogen-containing gases, *Nature*, 346, 552–554, doi:10.1038/346552a0, 1990.
- Lohmann, U. and Feichter, J.: Global indirect aerosol effects: a review, *Atmos. Chem. Phys.*, 5, 715–737, doi:10.5194/acp-5-715-2005, 2005.
- Loukonen, V., Kurtén, T., Ortega, I. K., Vehkamäki, H., Pádua, A. A. H., Sellegri, K., and Kulmala, M.: Enhancing effect of dimethylamine in sulfuric acid nucleation in the presence of water – a computational study, *Atmos. Chem. Phys.*, 10, 4961–4974, doi:10.5194/acp-10-4961-2010, 2010.
- Loukonen, V., Kuo, I. F. W., McGrath, M. J., and Vehkamäki, H.: On the stability and dynamics of (sulfuric acid)(ammonia) and (sulfuric acid)(dimethylamine) clusters: A first-principles molecular dynamics investigation, *Chem. Phys.*, 428, 164–174, doi:10.1016/j.chemphys.2013.11.014, 2014.
- Lunn, F. and Van de Vyver, J.: Sampling and analysis of air in pig houses, *Agric. Environ.*, 3, 159–169, doi:10.1016/0304-1131(77)90007-8, 1977.
- Mader, B. T., Yu, J. Z., Xu, J. H., Li, Q. F., Wu, W. S., Flagan, R. C., and Seinfeld, J. H.: Molecular composition of the water-soluble fraction of atmospheric carbonaceous aerosols collected during ACE-Asia, *J. Geophys. Res.*, 109, 1–13, doi:10.1029/2003JD004105, 2004.
- Müller, C., Iinuma, Y., Karstensen, J., van Pinxteren, D., Lehmann, S., Gnauk, T., and Herrmann, H.: Seasonal variation of aliphatic amines in marine sub-micrometer particles at the Cape Verde islands, *Atmos. Chem. Phys.*, 9, 9587–9597, doi:10.5194/acp-9-9587-2009, 2009.
- Murphy, S. M., Sorooshian, A., Kroll, J. H., Ng, N. L., Chhabra, P., Tong, C., Surratt, J. D., Knipping, E., Flagan, R. C., and Seinfeld, J. H.: Secondary aerosol formation from atmospheric reactions of aliphatic amines, *Atmos. Chem. Phys.*, 7, 2313–2337, doi:10.5194/acp-7-2313-2007, 2007.
- Myhre, G., Shindell, D., Bréon, F.-M., Collins, W., Fuglestedt, J., Huang, J., Koch, D., Lamarque, J.-F., Lee, D., Mendoza, B., Nakajima, T., Robock, A., Stephens, G., Takemura, T., and Zhang, H.: Anthropogenic and Natural Radiative Forcing, in: *Climate Change 2013: The Physical Science Basis. Contribution of Working Group I to the Fifth Assessment Report of the Intergovernmental Panel on Climate Change*, Cambridge, United Kingdom and New York, NY, USA, 81 pp., 2013.
- Nadykto, A., Herb, J., Yu, F., Xu, Y., and Nazarenko, E. S.: Estimating the lower limit of the impact of amines on nucleation in the Earth's atmosphere, *Entropy*, 17, 2764–2780, doi:10.3390/e17052764, 2015.
- Ortega, I. K., Kupiainen, O., Kurtén, T., Olenius, T., Wilkman, O., McGrath, M. J., Loukonen, V., and Vehkamäki, H.: From quantum chemical formation free energies to evaporation rates, *Atmos. Chem. Phys.*, 12, 225–235, doi:10.5194/acp-12-225-2012, 2012.
- Ortega, I. K., Donahue, N. M., Kurtén, T., Kulmala, M., Focsa, C., and Vehkamäki, H.: Can highly oxidized organics contribute to atmospheric new particle formation?, *J. Phys. Chem. A*, 120, 1452–1458, doi:10.1021/acs.jpca.5b07427, 2016.
- Pósfai, M., Simonics, R., Li, J., Hobbs, P. V., and Buseck, P. R.: Individual aerosol particles from biomass burning in southern Africa: 1. Compositions and size distributions of carbonaceous particles, *J. Geophys. Res.-Atmos.*, 108, doi:10.1029/2002JD002291, 2003.
- Possanzini, M. and Di Palo, V.: Improved HPLC determination of aliphatic amines in air by diffusion and derivatization techniques, *Chromatographia*, 29, 151–154, doi:10.1007/BF02268702, 1990.
- Qiu, C. and Zhang, R.: Multiphase chemistry of atmospheric amines, *Phys. Chem. Phys.*, 15, 5738–5752, doi:10.1039/c3cp43446j, 2013.
- Qiu, C., Wang, L., Lal, V., Khalizov, A. F., and Zhang, R.: Heterogeneous reactions of alkylamines with ammonium sulfate and ammonium bisulfate, *Environ. Sci. Technol.*, 45, 4748–4755, doi:10.1021/es1043112, 2011.
- Rabaud, N. E., Ebeler, S. E., Ashbaugh, L. L., and Flocchini, R. G.: Characterization and quantification of odorous and non-odorous volatile organic compounds near a commercial dairy in California, *Atmos. Environ.*, 37, 933–940, doi:10.1016/S1352-2310(02)00970-6, 2003.
- Rey, M. A. and Pohl, C. A.: Novel cation column for separation of amines and six common inorganic cations, *J. Chrom.*, 739, 87–97, 1996.
- Rey, M. and Pohl, C.: Novel cation-exchange column for the separation of hydrophobic and/or polyvalent amines, *J. Chromatogr. A*, 997, 199–206, doi:10.1016/S0021-9673(03)00113-4, 2003.
- Rogge, W. F., Hildemann, L. M., Mazurek, M. A., Cass, G. R., and Simonelt, B. R. T.: Sources of fine organic aerosol. 1. Charbroilers and meat cooking operations, *Environ. Sci. Technol.*, 25, 1112–1125, doi:10.1021/es00018a015, 1991.
- Ruiz-Jimenez, J., Parshintsev, J., Laitinen, T., Hartonen, K., Petäjä, T., Kulmala, M., and Riekkola, M. L.: Influence of the sampling site, the season of the year, the particle size and the number of nucleation events on the chemical composition of atmospheric ultrafine and total suspended particles, *Atmos. Environ.*, 49, 60–68, doi:10.1016/j.atmosenv.2011.12.032, 2012.
- Saleh, R., Robinson, E. S., Tkacik, D. S., Ahern, A. T., Liu, S., Aiken, A. C., Sullivan, R. C., Presto, A. A., Dubey, M. K., Yokelson, R. J., Donahue, N. M., and Robinson, A. L.: Brownness of organics in aerosols from biomass burning linked to their black carbon content, *Nat. Geosci.*, 7, 647–650, doi:10.1038/ngeo2220, 2014.

- Schade, G. W. and Crutzen, P. J.: Emission of aliphatic amines from animal husbandry and their reactions: Potential source of N₂O and HCN, *J. Atmos. Chem.*, 22, 319–346, doi:10.1007/BF00696641, 1995.
- Schauer, J. J., Kleeman, M. J., Cass, G. R., and Simoneit, B. R. T.: Measurement of emissions from air pollution sources. 1. C1 through C29 organic compounds from meat charbroiling, *Environ. Sci. Technol.*, 33, 1566–1577, doi:10.1021/es980076j, 1999.
- Schmeltz, I., and Hoffmann, D.: Nitrogen-containing compounds in tobacco and tobacco smoke, *Chem. Rev.*, 77, 295–311, doi:10.1021/cr60307a001, 1977.
- Seo, S.-G., Ma, Z.-K., Jeon, J.-M., Jung, S.-C., and Lee, W.-B.: Measurements of key offensive odorants in a fishery industrial complex in Korea, *Atmos. Environ.*, 45, 2929–2936, doi:10.1016/j.atmosenv.2011.01.032, 2011.
- Silva, P. J., Erupe, M. E., Price, D., Elias, J., Malloy, Q. G. J., Li, Q., Warren, B., and Cocker, D. R.: Trimethylamine as precursor to secondary organic aerosol formation via nitrate radical reaction in the atmosphere, *Environ. Sci. Technol.*, 42, 4689–4696, doi:10.1021/es703016v, 2008.
- Smith, J. N., Dunn, M. J., VanReken, T. M., Iida, K., Stolzenburg, M. R., McMurry, P. H., and Huey, L. G.: Chemical composition of atmospheric nanoparticles formed from nucleation in Tecamac, Mexico: Evidence for an important role for organic species in nanoparticle growth, *Geophys. Res. Lett.*, 35, 2–6, doi:10.1029/2007GL032523, 2008.
- Smith, J. N., Barsanti, K. C., Friedli, H. R., Ehn, M., Kulmala, M., Collins, D. R., Scheckman, J. H., Williams, B. J., and McMurry, P. H.: Observations of aminium salts in atmospheric nanoparticles and possible climatic implications, *P. Natl. Acad. Sci. USA*, 107, 6634–6639, doi:10.1073/pnas.0912127107, 2010.
- Sobanska, S., Hwang, H., Choël, M., Jung, H.-J., Eom, H.-J., Kim, H., Barbillat, J., and Ro, C. U.: Investigation of the chemical mixing state of individual asian dust particles by the combined use of electron probe X-ray microanalysis and raman microspectrometry, *Anal. Chem.*, 84, 3145–3154, doi:10.1021/ac2029584, 2012.
- Sorooshian, A., Murphy, S. M., Hersey, S., Gates, H., Padró, L. T., Nenes, A., Brechtel, F. J., Jonsson, H., Flagan, R. C., and Seinfeld, J. H.: Comprehensive airborne characterization of aerosol from a major bovine source, *Atmos. Chem. Phys.*, 8, 5489–5520, doi:10.5194/acp-8-5489-2008, 2008.
- Sorooshian, A., Padro, L. T., Nenes, A., Feingold, G., McComiskey, A., Hersey, S. P., Gates, H., Jonsson, H. H., Miller, S. D., Stephens, G. L., Flagan, R. C., and Seinfeld, J. H.: On the link between ocean biota emissions, aerosol, and maritime clouds: Airborne, ground, and satellite measurements off the coast of California, *Global Biogeochem. Cy.*, 23, 1–15, doi:10.1029/2009GB003464, 2009.
- Stephen and et al.: A comparative risk assessment of burden of disease and injury attributable to 67 risk factors and risk factor clusters in 21 regions, 1990–2010: a systematic analysis for the Global Burden of Disease Study 2010, *Lancet*, 380, 2224–2260, doi:10.1016/S0140-6736(12)61766-8, 2012.
- Sun, Y., Zhuang, G., Tang, A., Wang, Y., and An, Z.: Chemical Characteristics of PM_{2.5} and PM₁₀ in Haze-Fog Episodes in Beijing, *Environ. Sci. Technol.*, 40, 3148–3155, doi:10.1021/es051533g, 2006.
- Suzuki, Y., Kawakami, M., and Akasaka, K.: ¹H NMR application for characterizing water-soluble organic compounds in urban atmospheric particles, *Environ. Sci. Technol.*, 35, 2656–2664, doi:10.1021/es001861a, 2001.
- Takahama, S., Schwartz, R. E., Russell, L. M., Macdonald, A. M., Sharma, S., and Leaitch, W. R.: Organic functional groups in aerosol particles from burning and non-burning forest emissions at a high-elevation mountain site, *Atmos. Chem. Phys.*, 11, 6367–6386, doi:10.5194/acp-11-6367-2011, 2011.
- Tao, Y., Ye, X., Jiang, S., Yang, X., Chen, J., Xie, Y., and Wang, R.: Effects of amines on particle growth observed in new particle formation events, *J. Geos. Res.-Atmos.*, 121, 1–12, doi:10.1002/2015JD024245, 2015.
- Tröstl, J., Chuang, W. K., Gordon, H., Heinritzi, M., Yan, C., Molteni, U., Ahlm, L., Frege, C., Bianchi, F., Wagner, R., Simon, M., Lehtipalo, K., Williamson, C., Craven, J. S., Duplissy, J., Adamov, A., Almeida, J., Flagan, R. C., Franchin, A., Fuchs, C., Guida, R., Gysel, M., Pet, T., and Steiner, G.: The role of low-volatility organic compounds for initial particle growth in the atmosphere, *Nature*, 533, 1–56, doi:10.1038/nature18271, 2016.
- VandenBoer, T. C., Petroff, A., Markovic, M. Z., and Murphy, J. G.: Size distribution of alkyl amines in continental particulate matter and their online detection in the gas and particle phase, *Atmos. Chem. Phys.*, 11, 4319–4332, doi:10.5194/acp-11-4319-2011, 2011.
- VandenBoer, T. C., Markovic, M. Z., Petroff, A., Czar, M. F., Borduas, N., and Murphy, J. G.: Ion chromatographic separation and quantitation of alkyl methylamines and ethylamines in atmospheric gas and particulate matter using preconcentration and suppressed conductivity detection, *J. Chromatogr. A*, 1252, 74–83, doi:10.1016/j.chroma.2012.06.062, 2012.
- Van Neste, A., Duce, R. A., and Lee, C.: Methylamines in the marine atmosphere, *Geophys. Res. Lett.*, 14, 711–714, 1987.
- van Pinxteren, M., Fiedler, B., van Pinxteren, D., Iinuma, Y., Körtzinger, A., and Herrmann, H.: Chemical characterization of sub-micrometer aerosol particles in the tropical Atlantic Ocean: marine and biomass burning influences, *J. Atmos. Chem.*, 72, 105–125, doi:10.1007/s10874-015-9307-3, 2015.
- Verrielle, M., Plaisance, H., Depelchin, L., Benchabane, S., Locoge, N., and Meunier, G.: Determination of 14 amines in air samples using midjet impingers sampling followed by analysis with ion chromatography in tandem with mass spectrometry, *J. Environ. Monit.*, 14, 402–408, doi:10.1039/c2em10636a, 2012.
- Wang, L., Khalizov, A. F., Zheng, J., Xu, W., Ma, Y., Lal, V., and Zhang, R.: Atmospheric nanoparticles formed from heterogeneous reactions of organics, *Nat. Geosci.*, 3, 238–242, doi:10.1038/ngeo778, 2010a.
- Wang, L., Vinita, L., Khalizov, A. F., and Zhang, R.: Heterogeneous chemistry of alkylamines with sulfuric acid: Implications for atmospheric formation of alkylammonium sulfates, *Environ. Sci. Technol.*, 44, 2461–2465, doi:10.1021/es9036868, 2010b.
- Willis, M. D., Burkart, J., Thomas, J. L., Köllner, F., Schneider, J., Bozem, H., Hoor, P. M., Aliabadi, A. A., Schulz, H., Herber, A. B., Leaitch, W. R., and Abbatt, J. P. D.: Growth of nucleation mode particles in the summertime Arctic: a case study, *Atmos. Chem. Phys.*, 16, 7663–7679, doi:10.5194/acp-16-7663-2016, 2016.
- Yang, H., Xu, J., Wu, W. S., Wan, C. H., and Yu, J. Z.: Chemical characterization of water-soluble organic aerosols at Jeju Island collected during ACE-Asia, *Environ. Chem.*, 1, 13–17, doi:10.1071/EN04006, 2004.

- Yang, H., Jian, Z. Y., Hang Ho, S. S., Xu, J., Wu, W. S., Chun, H. W., Wang, X., Wang, X., and Wang, L.: The chemical composition of inorganic and carbonaceous materials in PM_{2.5} in Nanjing, China, *Atmos. Environ.*, 39, 3735–3749, doi:10.1016/j.atmosenv.2005.03.010, 2005.
- Yao, L., Wang, M.-Y., Wang, X.-K., Liu, Y.-J., Chen, H.-F., Zheng, J., Nie, W., Ding, A.-J., Geng, F.-H., Wang, D.-F., Chen, J.-M., Worsnop, D. R., and Wang, L.: Detection of atmospheric gaseous amines and amides by a high-resolution time-of-flight chemical ionization mass spectrometer with protonated ethanol reagent ions, *Atmos. Chem. Phys.*, 16, 14527–14543, doi:10.5194/acp-16-14527-2016, 2016.
- Youn, J.-S., Crosbie, E., Maudlin, L. C., Wang, Z., and Sorooshian, A.: Dimethylamine as a major alkyl amine species in particles and cloud water: Observations in semi-arid and coastal regions, *Atmos. Environ.*, 122, 250–258, doi:10.1016/j.atmosenv.2015.09.061, 2015.
- Yu, H., McGraw, R., and Lee, S. H.: Effects of amines on formation of sub-3 nm particles and their subsequent growth, *Geophys. Res. Lett.*, 39, 2–7, doi:10.1029/2011GL050099, 2012.
- Zauscher, M. D., Wang, Y., Moore, M. J. K., Gaston, C. J., and Prather, K. A.: Air quality impact and physicochemical aging of biomass burning aerosols during the 2007 San Diego wildfires, *Environ. Sci. Technol.*, 47, 7633–7643, doi:10.1021/es4004137, 2013.
- Zhang, Q. and Anastasio, C.: Free and combined amino compounds in atmospheric fine particles (PM_{2.5}) and fog waters from Northern California, *Atmos. Environ.*, 37, 2247–2258, doi:10.1016/S1352-2310(03)00127-4, 2003.
- Zhao, J., Smith, J. N., Eisele, F. L., Chen, M., Kuang, C., and McMurry, P. H.: Observation of neutral sulfuric acid-amine containing clusters in laboratory and ambient measurements, *Atmos. Chem. Phys.*, 11, 10823–10836, doi:10.5194/acp-11-10823-2011, 2011.
- Zollner, J. H., Glasoe, W. A., Panta, B., Carlson, K. K., McMurry, P. H., and Hanson, D. R.: Sulfuric acid nucleation: power dependencies, variation with relative humidity, and effect of bases, *Atmos. Chem. Phys.*, 12, 4399–4411, doi:10.5194/acp-12-4399-2012, 2012.